

Quantum Numbers

- Used to _____ an _____ in an _____

n

- Represents _____ energy level of _____
- _____ of _____ in an _____

Example: What is the maximum number of electrons that can be in the _____ main energy level?

l

- The _____
- Describes the _____ within an _____
- _____ of orbital _____ possible is _____

Orbital Shapes

designated _____

- level 1: _____
- level 2: _____
- level 3: _____
- level 4: _____

How many electrons can each sublevel hold?

- s = 1 orbital \times 2 e⁻/orbital = _____e⁻
- p = 3 orbitals \times 2 e⁻/orbital = _____e⁻
- d = 5 orbitals \times 2 e⁻/orbital = _____e⁻
- f = 7 orbitals \times 2 e⁻/orbital = _____e⁻

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