

A. Identify species which the  $K_{eq}$  is associated to. Part of this is to be able to identify which species are important.



Classroom:  $K_{eq}$  is a ratio of products where the reactants are in the denominator. Therefore there must be a right side species which must be a product. Therefore it is important.

A. Equilibrium:

• If the  $K_{eq}$  is a constant, but if the reactants are not in the denominator, then it is not a constant.

• If the concentration of reactants is not in the denominator, then it is not a constant. If the concentration of the products is not in the denominator, then it is not a constant.



$$K_{eq} = \frac{[C][D]}{[A][B]}$$

Ex:  $K_{eq}$  is \_\_\_\_\_ reaction is favored

$K_{eq} > 1$	Products
$K_{eq} < 1$	Reactants
$K_{eq} = 1$	Equilibrium

Ex: 10% Calculate  $K_{eq}$  for the following reaction using the given equilibrium concentrations. Then determine whether the forward or reverse reaction is favored.



$$\frac{[H_2O_{(g)}]}{[H_2O_{(l)}]}$$

$$\frac{[0.001]}{[0.001]}$$

$$K_{eq} = 1$$

# Note Taking Guide Episode 1202 Answers

**Xiang Xie**



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