

Background

- 1. Study of the \_\_\_\_\_ multi \_\_\_\_\_ relationships in a \_\_\_\_\_ structure \_\_\_\_\_
- 2. Research on \_\_\_\_\_ balanced \_\_\_\_\_ equations

2.  $4\text{Hg} + \text{Cl}_2 \rightarrow 2\text{Hg}_2\text{Cl}_2$

2. Magnesium reacting with a single oxygen atom produces 2 formula units of magnesium oxide

The \_\_\_\_\_ in a \_\_\_\_\_ reaction \_\_\_\_\_ give the \_\_\_\_\_ ratio \_\_\_\_\_ for the \_\_\_\_\_ involved in the reaction \_\_\_\_\_

E.g. Problem

When aluminum aluminum reacts with elemental oxygen, aluminum oxide is produced



molar ratios

4	Al	3	O <sub>2</sub>
2	Al <sub>2</sub>	3	Al <sub>2</sub> O <sub>3</sub>
2	O <sub>2</sub>	3	Al <sub>2</sub> O <sub>3</sub>

2. If you start with 4 moles of Al, how many moles of Al<sub>2</sub>O<sub>3</sub> will be produced?

2 moles Al (moles Al<sub>2</sub>O<sub>3</sub>) and 3 moles O<sub>2</sub> = 2 moles Al<sub>2</sub>O<sub>3</sub>

# Note Taking Guide Episode 801 Answers

**Daniel F McAuley**



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