

Atomic Number

- Increase across a period
- increase down a family

Atomic Mass

- increase across a period
- increase down a family

Atomic Radius

half the distance between nuclei of two like atoms.

- decreases across a period (due to increased attraction of more positive nuclei to electron clouds)
- increase down a family (due to addition of an energy level)

Electron Affinity

amount of energy when an atom gains an electron

- increases across a period
- decreases down a family

Ionization Energy

amount of energy required to remove an electron from a neutral atom

- increase across a period
- decrease down a family

Electron Distribution - Forming Ions

- Atoms become ions by gaining or losing electrons to become an ion.
- An atom is stable when it has a full outer energy level (8 valence electrons).
- Atoms that lose (give away) electrons become positive charged ions. Atoms that gain (steal) electrons become negative charged ions.

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Table of Contents Note Taking Guide Chemistry Episode 403

1. Understanding the eBook Note Taking Guide Chemistry Episode 403
 - The Rise of Digital Reading Note Taking Guide Chemistry Episode 403
 - Advantages of eBooks Over Traditional Books
2. Identifying Note Taking Guide Chemistry Episode 403
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in a Note Taking Guide Chemistry Episode 403
 - User-Friendly Interface
4. Exploring eBook Recommendations from Note Taking Guide Chemistry Episode 403
 - Personalized Recommendations
 - Note Taking Guide Chemistry Episode 403 User Reviews and Ratings
 - Note Taking Guide Chemistry Episode 403 and Bestseller Lists
5. Accessing Note Taking Guide Chemistry Episode 403 Free and Paid eBooks
 - Note Taking Guide Chemistry Episode 403 Public Domain eBooks
 - Note Taking Guide Chemistry Episode 403 eBook Subscription Services
 - Note Taking Guide Chemistry Episode 403 Budget-Friendly Options

6. Navigating Note Taking Guide Chemistry Episode 403 eBook Formats
 - ePub, PDF, MOBI, and More
 - Note Taking Guide Chemistry Episode 403 Compatibility with Devices
 - Note Taking Guide Chemistry Episode 403 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Note Taking Guide Chemistry Episode 403
 - Highlighting and Note-Taking Note Taking Guide Chemistry Episode 403
 - Interactive Elements Note Taking Guide Chemistry Episode 403
8. Staying Engaged with Note Taking Guide Chemistry Episode 403
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Note Taking Guide Chemistry Episode 403
9. Balancing eBooks and Physical Books Note Taking Guide Chemistry Episode 403
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Note Taking Guide Chemistry Episode 403
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Note Taking Guide Chemistry Episode 403
 - Setting Reading Goals Note Taking Guide Chemistry Episode 403
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Note Taking Guide Chemistry Episode 403
 - Fact-Checking eBook Content of Note Taking Guide Chemistry Episode 403
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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