

Introduction:

The laboratory method used in the experiment is titration. Titration is a method used in measuring the amount of an analytical reagent necessary to react quantitatively with the sample. Acid-base titrations are important for counting concentrations of acids and bases. The point at which the equivalence point has been reached is called the theoretical end point. The titration is detected by some physical change, such as colour change, to indicate the end point. The indicator solution used in this experiment is phenolphthalein which is a colourless in acidic solution. When the reaction reached the endpoint, the solution becomes pale pink. One or two drops of phenolphthalein are sufficient for the titration. Addition of too much indicator will necessitate an indicator blank. The indicator should give a clear visual change in the liquid being titrated once the reaction between the standard solution and the substance is practically complete. The indicators can change colour because their ions have colours that are different from the undissociated molecule (Tan, 2010). The acid and base titrations used the Arrhenius theory. This theory stated that acid are substance which produce hydrogen ions in solution and bases are substance which produce hydroxide ion in solution. In titration, standard solutions are now usually expressed in terms of molarity, M.

Molarity, M = moles of solute/volume of solutions in litres

Objective :

One of the purposes of this experiment is to demonstrate the basic laboratory technique of titration towards the students. This experiment enables a hand-on practice done by the students themselves. Thus, the students will know how to handle the materials, apparatus and the correct way to conduct the experiment as well as being careful to avoid any experimental error. Moreover, this experiment also teach the students on the correct steps to calculate percentage of substance, molarity of solution and other calculations involve throughout the experiment.

Apparatus :

Burette with stand
Pipette
1-L volumetric flask with stopper
250 mL Erlenmeyer flasks
Retort stand with clamp

Reagents :

Sodium hydroxide pellets
Potassium hydrogen phthalate (KHP)
Phenolphthalein
Unknown vinegar

Sample Introduction Lab Report

Vicky McMillan

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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How to Write a Lab Report Jerome N. Borowick, 2000 This guide outlines an effective methodology for writing the experimental laboratory report showing how skills that emphasize correct grammar and appropriate style must be adapted to writing reports with a purpose reports that emphasize structure and content to persuade the readers It first covers basic principles then explores each section of a report step by step with sample report sections and critiques The Laboratory Report Writing Process Principles of Clear Lab Report Writing Rules of Practice for Lab Report Writing Graphics The Title Page and Table of Contents The Beginning of the Report The Body of the Report The Ending of the Report A Sample Student Lab Report For anyone who must write lab reports as part of their professional responsibilities **Successful Lab Reports** Christopher S. Lobban, MarLa Schefter, 1992-02-28 Shows science students how to write a clear and to the point laboratory report **Laboratory Tests on Samples from Borrow Area D (appendix 2 Added), Keene Creek Dam, Rogue River Project, Oregon**, 1957

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1. This book is structured into several chapters, namely:
 - Chapter 1: Introduction to Sample Introduction Lab Report
 - Chapter 2: Essential Elements of Sample Introduction Lab Report
 - Chapter 3: Sample Introduction Lab Report in Everyday Life
 - Chapter 4: Sample Introduction Lab Report in Specific Contexts
 - Chapter 5: Conclusion
2. In chapter 1, this book will provide an overview of Sample Introduction Lab Report. This chapter will explore what Sample Introduction Lab Report is, why Sample Introduction Lab Report is vital, and how to effectively learn about Sample Introduction Lab Report.
3. In chapter 2, the author will delve into the foundational concepts of Sample Introduction Lab Report. The second chapter will elucidate the essential principles that need to be understood to grasp Sample Introduction Lab Report in its entirety.
4. In chapter 3, this book will examine the practical applications of Sample Introduction Lab Report in daily life. The third chapter will showcase real-world examples of how Sample Introduction Lab Report can be effectively utilized in everyday scenarios.
5. In chapter 4, this book will scrutinize the relevance of Sample Introduction Lab Report in specific contexts. This chapter will explore how Sample Introduction Lab Report is applied in specialized fields, such as education, business, and technology.
6. In chapter 5, the author will draw a conclusion about Sample Introduction Lab Report. The final chapter will summarize the key points that have been discussed throughout the book.

This book is crafted in an easy-to-understand language and is complemented by engaging illustrations. It is highly recommended for anyone seeking to gain a comprehensive understanding of Sample Introduction Lab Report.

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