

## Chapter 12 Review 1: Covalent Bonds and Molecular Structure

1) How are ionic bonds and covalent bonds different, and what types of elements combine to form each?

**Ionic bonds result from the transfer of electrons from one atom to another (formed by a metal and a non-metal).**

**Covalent bonds result from two atoms sharing electrons (formed by 2 or more non-metals).**

2) How are nonpolar covalent bonds different from covalent bonds, and what types of elements combine to form each?

**polar – electrons are shared unequally – between nonmetals with different electronegativities**

**nonpolar – electrons are shared equally – between multiple atoms of the same element**

3) Identify the type(s) of bond(s) found in the following molecules:

- a.  $\text{CCl}_4$  \_\_\_\_\_ covalent \_\_\_\_\_
- b.  $\text{Li}_2\text{O}$  \_\_\_\_\_ ionic \_\_\_\_\_
- c.  $\text{NF}_3$  \_\_\_\_\_ covalent \_\_\_\_\_
- d.  $\text{CaSO}_4$  \_\_\_\_\_ ionic and covalent \_\_\_\_\_
- e.  $\text{SO}_2$  \_\_\_\_\_ covalent \_\_\_\_\_
- f.  $\text{Mg(OH)}_2$  \_\_\_\_\_ ionic and covalent \_\_\_\_\_

4) Define electronegativity.

**Electronegativity is the ability for an atom in a molecule to attract electrons to itself.**

5) Use electronegativity values to place the following elements in increasing order: F, N, Si, C, O

**$\text{Si} < \text{C} < \text{N} < \text{O} < \text{F}$**

6) Determine if the bond between atoms in each example below is nonpolar covalent, polar covalent, or ionic.

- |                   |                          |                  |                          |
|-------------------|--------------------------|------------------|--------------------------|
| a. $\text{H}_2$   | _____ <u>npv</u> _____   | e. $\text{NF}_3$ | _____ <u>pc</u> _____    |
| b. $\text{PCl}_3$ | _____ <u>pc</u> _____    | f. $\text{MgO}$  | _____ <u>ionic</u> _____ |
| c. $\text{F}_2$   | _____ <u>npv</u> _____   | g. $\text{CH}_4$ | _____ <u>npv</u> _____   |
| d. $\text{NaBr}$  | _____ <u>ionic</u> _____ | h. $\text{HCl}$  | _____ <u>pc</u> _____    |

# Reviewing Key Concepts Chapter 12

**PT Brinkman**



## Reviewing Key Concepts Chapter 12:

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**Miscellaneous Publication**, 1978

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The novel begins with Richard Papen, a bright but troubled young man, arriving at Hampden College. Richard is immediately drawn to the group of students who call themselves the Classics Club. The club is led by Henry Winter, a brilliant and charismatic young man. Henry is obsessed with Greek mythology and philosophy, and he quickly draws Richard into his world. The other members of the Classics Club are equally as fascinating. Bunny Corcoran is a wealthy and spoiled young man who is always looking for a good time. Charles Tavis is a quiet and reserved young man who is deeply in love with Henry. Camilla Macaulay is a beautiful and intelligent young woman who is drawn to the power and danger of the Classics Club. The students are all deeply in love with Morrow, and they are willing to do anything to please him. Morrow is a complex and mysterious figure, and he seems to be manipulating the students for his own purposes. As the students become more involved with Morrow, they begin to commit increasingly dangerous acts.

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