

Stoichiometry: Practice Assessment

Name: _____

5&6B - Using a balanced equation, given the grams of reactant, calc. the grams of product



Balance the equation provided. Calculate the number of moles of sodium bromide that would be produced if 25.5 moles of gallium (III) bromide are used (with an excess of sodium sulfite)

$$25.5\text{mol GaBr}_3 \times 6\text{mol NaBr}/2\text{mol GaBr}_3 = 75.75\text{mol NaBr}$$

2) Tin (II) oxide reacts with nitrogen trifluoride to produce tin (II) fluoride and dinitrogen trioxide. Using the following reaction, calculate the grams produced of each product from 50.0 grams of nitrogen trifluoride.



$$50\text{g NF}_3 \times 1\text{mol NF}_3 / 71.001\text{g NF}_3 = 0.704\text{mol NF}_3 \times 2\text{mol SnF}_2 / 2\text{mol NF}_3 \times 156.69\text{g SnF}_2 / 1\text{mol SnF}_2 = 165.5\text{g SnF}_2$$

SnF₂

$$71.001\text{g NF}_3 = 0.704\text{mol NF}_3 \times 1\text{mol N}_2\text{O}_3 / 2\text{mol NF}_3 \times 76.014\text{g N}_2\text{O}_3 / 1\text{mol N}_2\text{O}_3 =$$

26.75g N₂O₃

3) A piece of copper wire is placed in silver nitrate. Silver and aqueous copper (II) nitrate is produced. The initial mass of the copper before the reaction was 2.74g and the mass remaining after the reaction was 2.32g. What is the mass of silver produced?



$$2.74\text{g} - 2.32\text{g} = 0.42\text{g}$$

$$0.42\text{g Cu} \times 1\text{mol Cu} / 63.55\text{g Cu} \times 2\text{mol Ag} / 1\text{mol Cu} \times 107.87\text{g Ag} / 1\text{mol Ag} = 1.43\text{g Ag}$$

5&6B - After calculating theoretical yield in grams, calculate the percent yield. Given experimental data, determine the limiting reactant and the excess reactant

4) Through experimentation, you observe the reaction of iron (III) phosphate with sodium sulfate to make iron (III) sulfate and sodium phosphate. You began the experiment with 25g of iron (III) phosphate and 182g of sodium sulfate. How many grams of iron (III) sulfate should you have been able to produce? What is the percent yield if 18.5g of iron (III) sulfate are actually made?



$$25\text{g FePO}_4 \times 1\text{mol FePO}_4 / 159.815\text{g FePO}_4 = 0.156\text{mol FePO}_4 / 2\text{mol per rxn} = 0.083\text{rxns (LR)}$$

$$182\text{g Na}_2\text{SO}_4 \times 1\text{mol Na}_2\text{SO}_4 / 142.04\text{g Na}_2\text{SO}_4 = 1.28\text{mol Na}_2\text{SO}_4 / 3\text{mol per rxn} = 0.427\text{rxns}$$

$$0.083\text{mol FePO}_4 \times 1\text{mol Fe}_2(\text{SO}_4)_3 / 2\text{mol FePO}_4 \times 399.873\text{g Fe}_2(\text{SO}_4)_3 / 1\text{mol Fe}_2(\text{SO}_4)_3 = 33.19\text{g}$$

$$(18.5\text{g} / 33.19\text{g}) \times 100 = 56\%$$

Review Stoichiometry Section 2 Answer Key

Ensheng Dong



Review Stoichiometry Section 2 Answer Key:

This book delves into Review Stoichiometry Section 2 Answer Key. Review Stoichiometry Section 2 Answer Key is a crucial topic that needs to be grasped by everyone, from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Review Stoichiometry Section 2 Answer Key, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:

- Chapter 1: Introduction to Review Stoichiometry Section 2 Answer Key
- Chapter 2: Essential Elements of Review Stoichiometry Section 2 Answer Key
- Chapter 3: Review Stoichiometry Section 2 Answer Key in Everyday Life
- Chapter 4: Review Stoichiometry Section 2 Answer Key in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, this book will provide an overview of Review Stoichiometry Section 2 Answer Key. This chapter will explore what Review Stoichiometry Section 2 Answer Key is, why Review Stoichiometry Section 2 Answer Key is vital, and how to effectively learn about Review Stoichiometry Section 2 Answer Key.

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4. In chapter 3, the author will examine the practical applications of Review Stoichiometry Section 2 Answer Key in daily life. This chapter will showcase real-world examples of how Review Stoichiometry Section 2 Answer Key can be effectively utilized in everyday scenarios.

5. In chapter 4, the author will scrutinize the relevance of Review Stoichiometry Section 2 Answer Key in specific contexts. This chapter will explore how Review Stoichiometry Section 2 Answer Key is applied in specialized fields, such as education, business, and technology.

6. In chapter 5, the author will draw a conclusion about Review Stoichiometry Section 2 Answer Key. The final chapter will summarize the key points that have been discussed throughout the book.

The book is crafted in an easy-to-understand language and is complemented by engaging illustrations. It is highly recommended for anyone seeking to gain a comprehensive understanding of Review Stoichiometry Section 2 Answer Key.

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