GAS STOICHIOMETRY PROBLEMS WORKSHEET 1

1. Given the unbalanced decomposition reaction of baking soda:

2NaHCO_{3 (c)} + heat ===> Na₂CO_{3 (c)} + CO_{3 (g)} + HOH (g) 42.0 grams of

baking soda? What volumes of carbon dioxide and water are produced at STP?

The catalytic decomposition of hydrogen peroxide is:

2H2O2 ===> 2H2O 0 + O2 10

Balance the reaction. How many moles of water and oxygen are produced by the decomposition of 68.0 grams of hydrogen peroxide? How many molecules of water and oxygen are produced? How many grams of each product are formed?

62, 11,01 , 1 mal fine , 2 mil fine 1 to 02, 11,00 [36 02, 15,00] 62, 11,02 , 1 mal fine , 1 mil fine (43.96, 02)

 If 156.0 grams of potassium metal reacts with excess water, potassium hydroxide and hydrogen gas are formed? What volume of hydrogen gas, in liters,

Be sure to balance the reaction first.

2Km + 2H₂O_m → 2KOH_{mo} + H_{2m}

156 J K × 1 mil K × 1 mil K × 17.46 U - 449.70 C H₂

156 J K × 1 mil K × 2 mil K × 1 mil Koti + 89.39 L Koti

4. Determine the number of moles of carbon dioxide gas, water, and sodium chloride formed by the reaction of 42.0 grams of sodium bicarbonate (baking soda) reacting with excess hydrochloric acid. The reaction is:

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Guillaume Favre

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