

Fill in the blank in each sentence and/or write a word.

1. The theory that explains the behavior of gases on the molecular level is called the kinetic theory, which is based on assumptions related to theoretical gas behavior, except that we use atoms molecules.
2. Use the kinetic theory to explain why a helium-filled balloon expands when it is brought inside the warmer room from the cold outdoors.
 A helium-filled balloon is a pressure vessel, which has some helium gas. According to the kinetic theory, the gas molecules within the balloon have kinetic energy and are constantly in motion. When the balloon is brought inside the warmer room, the gas molecules gain kinetic energy and move faster, causing the balloon to expand.
3. Define "ideal gas."

A hypothetical gas that obeys the ideal gas law, which is a mathematical equation that relates the pressure, volume, and temperature of a gas.

4. Gases deviate most from ideal behavior under conditions of very low temperature and very high pressure.
5. Pressure is defined as force per unit area. The earth's atmosphere has weight, which creates atmospheric pressure.
6. Standard temperature has been established as 0 °C or 32 °F. Standard pressure has been established as 1.013 atm, 101.3 kPa, 760 mm Hg, or 14.7 psi. The standard unit for standard temperature and pressure is STP.
7. A manometer is used to measure atmospheric pressure, which is absolute, is used to measure gas pressure.
8. When atmospheric pressure increases, how does the height of the column of mercury change? Increases
9. At constant temperature, the volume of a sample of gas decreases as temperature increases.
10. At constant pressure, the volume of a sample of gas is directly proportional to a temperature increased as the temperature decreases.

Review Sheet Unit 9 Chemistry Answers

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