

10. Circle the element in each of the following pairs with the greater (1<sup>st</sup>) ionization energy:
- a) lithium or boron  
 b) magnesium or strontium  
 c) cesium or lead  
 d) sodium or argon
11. Arrange the following elements in order of increasing ionization energy:
- a) Sr, Be, Mg  
 b) Bi, Cs, Ba  
 c) Na, Al, S
12. Why is there a very large increase between the 1<sup>st</sup> and 2<sup>nd</sup> ionization energies for the elements in group 1?

Name: \_\_\_\_\_ Date: \_\_\_\_\_ Class period: \_\_\_\_\_



### Unit: Periodic Trends “Ionic Size and E.N. Trend” - WS #4

Directions: Please answer each fill in the blank with the best answer. ☺

- A cation is an ion with an overall \_\_\_\_\_ charge.
- Cations are made when an atom \_\_\_\_\_.
- The size of a positive ion is always \_\_\_\_\_ than the size of the neutral atom of the same element.
- As the positive charge on an ion increases, the size of the atom will \_\_\_\_\_.
- An anion is an ion with an overall \_\_\_\_\_ charge.
- Anions are made when at atom \_\_\_\_\_.
- The size of a negative ion is always \_\_\_\_\_ than the size of the neutral atom of the same element.
- As the negative charge on an ion increases, the size of the atom will \_\_\_\_\_.
- Which element/ion has the **biggest** atomic/ionic size in each of the following pairs? Please circle your answer.
 

a) Na or Na <sup>+</sup>	d) Al or Al <sup>3+</sup>
b) S or S <sup>2-</sup>	e) Cu <sup>+</sup> or Cu <sup>2+</sup>
c) O <sup>-</sup> or O <sup>2-</sup>	
- \_\_\_\_\_ is the tendency for the atoms of the element to attract electrons when they are chemically combined with another element.

# Periodic Trends Unit Exam And Answers

**Chao Zhang**



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