

# Graphing Periodic Trends Lab

Name \_\_\_\_\_ Hr \_\_\_\_\_

**PURPOSE:** To determine how certain properties are periodic when the elements are arranged in periods or groups.

**DISCUSSION:** There are many periodic trends that occur as you cross a period or go down a group/family on the Periodic Table. Graphing various elemental properties will help to illustrate these trends.

## REFERENCE:

Angstrom - A unit of distance equal to one ten-billionth of a meter.

Atomic Radius - The distance from the center of the nucleus to the valence electron energy level.

Ionization energy - The energy required to remove an electron from an atom.

## DATA

**Properties of the First 18 Elements**

| Element | Atomic Number | Atomic Radius (Angstroms) | Ionization Energy |
|---------|---------------|---------------------------|-------------------|
| H       | 1             | 0.79                      | 314               |
| He      | 2             | 0.49                      | 567               |
| Li      | 3             | 2.05                      | 124               |
| Be      | 4             | 1.4                       | 215               |
| B       | 5             | 1.17                      | 191               |
| C       | 6             | 0.91                      | 260               |
| N       | 7             | 0.75                      | 335               |
| O       | 8             | 0.65                      | 314               |
| F       | 9             | 0.57                      | 402               |
| Ne      | 10            | ?                         | ?                 |
| Na      | 11            | 2.23                      | 119               |
| Mg      | 12            | 1.72                      | 176               |
| Al      | 13            | 1.82                      | 138               |
| Si      | 14            | 1.46                      | 168               |
| P       | 15            | 1.23                      | 242               |
| S       | 16            | 1.09                      | 239               |
| Cl      | 17            | 0.97                      | 299               |
| Ar      | 18            | ?                         | ?                 |

Using the data above, make a line graph the following data. Draw a dark vertical line to separate the 3 periods. Give each graph a title and label both the x-axis and the y-axis. Connect your points. Predict the plot points for elements with question marks.

1. Atomic Radius (y) vs. Atomic Number (x) for the first 3 periods (18 elements – 3 graphs)
2. Ionization energy (y) vs. Atomic Number (x) for the first 3 periods (18 elements – 3 graphs)

# Periodic Trends Graphing Lab Answers

**R Bogdan**



## Periodic Trends Graphing Lab Answers:

*Computer Based Projects for a Chemistry Curriculum* Thomas J. Manning, Aurora P. Gramatges, 2013-04-04 This e book is a collection of exercises designed for students studying chemistry courses at a high school or undergraduate level The e book contains 24 chapters each containing various activities employing applications such as MS excel spreadsheets and Spartan computational modeling Each project is explained in a simple easy to understand manner The content within this book is suitable as a guide for both teachers and students and each chapter is supplemented with practice guidelines and exercises Computer Based Projects for a Chemistry Curriculum therefore serves to bring computer based learning a much needed addition in line with modern educational trends to the chemistry classroom

*General, Organic, and Biological Chemistry Study Guide and Selected Solutions* Karen C. Timberlake, 2001-11 Keyed to the learning goals in the text this guide is designed to promote active learning through a variety of exercises with answers and mastery exams The guide also contains complete solutions to odd numbered problems

**Laboratory Manual** Jo A. Beran, 1990

*The Science Teachers Bulletin*, 1989

**Chemistry** Karen Timberlake, 1999 Suitable for one or two term lab courses covering general organic and biological chemistry this new edition written by Karen Timberlake features many improvements to the insightful experiments that have made it the leading lab manual Each experiment encourages critical thinking with laboratory goals discussion of related concepts clear instructions new pre lab questions and comprehensive report pages Forty one experiments illustrate the basic principles of chemistry

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*CAS Review*, 1975

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*The World of Chemistry* Jones, Jones Staff, 1996

**Prentice Hall Science Explorer** Michael J. Padilla, 2002

**Software Reference Guide**, 1988

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**Physics Briefs**, 1987

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**Exemplary Programs in Introductory College Mathematics** Susan S. Lenker, 1998 This handbook contains a collection of the winning entries in the first INPUT Competition part of the INPUT Innovative Programs Using Technology Project The INPUT Project was designed to improve instruction by recognizing and rewarding college instructors who rethought the mathematical content of their introductory mathematics courses with innovative uses of technology The targeted introductory mathematics courses were developmental mathematics precalculus business mathematics and introductory statistics

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