

## Periodic Trends Worksheet

Use the periodic table and your knowledge of periodic trends to answer the following questions.

1. Which atom in each pair has the larger atomic radius?

a) Li or **K**      b) **Ca** or Ni      c) **Ga** or B      d) O or **C**      e) Cl or **Br**  
f) Be or **Ba**      g) **Si** or S      h) Fe or **Au**

2. What is the periodic trend for atomic size from top to bottom in a group? from left to right in a period?

Top to bottom: increases due to adding energy levels. Left to Right: decreases

3. Why do atoms get smaller as you move left to right in a period?

due to increased nuclear charge and no added shielding.

4. Which element in each pair has a larger ionization energy?

a) Na or **O**      b) **Be** or Ba      c) **Ar** or F      d) **Cu** or Ra      e) I or **Ne**  
f) K or **V**      g) **Ca** or Fr      h) W or **Se**

5. Explain the relationship between the relative size of an ion to its neutral atom and the charge on the ions.

Positive ions are smaller than the neutral atom it comes from, negative ions are larger than the neutral atom it comes from

6. Which particle has the larger radius in each atom/ion pair?

a) Na,  $\text{Na}^+$       b) S,  $\text{S}^{2-}$       c) I,  $\text{I}^-$       d) Al,  $\text{Al}^{3+}$

7. What is ionization energy? What is first ionization energy?

The amount of energy needed to remove an electron from an atom. First ionization energy is the energy required to remove a single valence electron from an atom.

8. What is the periodic trend for first ionization energy?

It increases going from left to right due to increased nuclear charge, and decreases going down a group due to increased distance from the nucleus and increased shielding.

9. Arrange the following groups of elements in order of increasing ionization energy.

a) Be, Mg, Sr      b) Bi, Cs, Ba      c) Na, Al, S  
high, middle, low      high, low, middle      low, middle, high

10. Which element in each pair has a higher electronegativity value?

a) Cl, **F**      b) C, **N**      c) **Mg**, Ne      d) **As**, Ca

11. What is the periodic trend for electronegativity? Increases going left to right, decreases going down.

# Periodic Trends Extension Questions Answers

**S Ben Porath**



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**Merrill Chemistry** Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998      Handbook of Nuclear Physics Isao Tanihata, Hiroshi Toki, Toshitaka Kajino, 2023-09-04 This handbook is a comprehensive systematic source of modern nuclear physics. It aims to summarize experimental and theoretical discoveries and an understanding of unstable nuclei and their exotic structures which were opened up by the development of radioactive ion RI beam in the late 1980s. The handbook comprises three major parts. In the first part the experiments and measured facts are well organized and reviewed. The second part summarizes recognized theories to explain the experimental facts introduced in the first part. Reflecting recent synergistic progress involving both experiment and theory the chapters both parts are mutually related. The last part focuses on cosmo nuclear physics one of the mainstream subjects in modern nuclear physics. Those comprehensive topics are presented concisely. Supported by introductory reviews all chapters are designed to present their topics in a manner accessible to readers at the graduate level. The book therefore serves as a valuable source for beginners as well helping them to learn modern nuclear physics.

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