

Lab 6: CHEMICAL EQUILIBRIUM LAB

Introduction: In this experiment you will start with a reaction that is at equilibrium. You will then change the concentrations of various ions that are present and note the effect on the state of equilibrium. The equilibrium reaction chosen has a colored product, so changes in its concentration can be readily detected by a corresponding change in color intensity.

Procedure:

1. Pour about 25 mL of 0.0020 M potassium thiocyanate solution, KSCN, and 25 mL of distilled water into a beaker.
 - a. What is the color of this solution? _____
 - b. What are the 2 ions in this solution? _____
2. Look at the bottle of 0.20M iron (III)nitrate solution, $\text{Fe}(\text{NO}_3)_3$ at your lab station.
 - a. What is the color of this solution? _____
 - b. What are the 2 ions in this solution? _____
3. Add 5 drops of the $\text{Fe}(\text{NO}_3)_3$ solution to the beaker containing the potassium thiocyanate solution, KSCN. Gently swirl the mixture. What is the color of this solution? _____
4. There are now 4 ions in this solution. What are the two combinations of ions that could account for the color in step three?

5. At your lab station there is a solution of potassium nitrate KNO_3 . What is the color of this solution? _____
What are the two ions in this solution? _____
6. Based on your observations in step 5 and your answer to question 4, what are the two ions that must be responsible for the color you observed in step 5?

Observing Chemical Equilibrium Lab 46 Answers

Daniela Niemeyer



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Chemistry Eugene LeMay, Jr., Herbert Beall, Karen M. Robblee, Douglas C. Brower, Prentice-Hall Staff, 2002-02

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The Iron(III) Thiocyanate Reaction Kevin C. de Berg, 2019-11-13 This Brief presents an historical investigation into the reaction between ferric ions and thiocyanate ions which has been viewed in different ways throughout the last two centuries Historically the reaction was used in chemical analysis and to highlight the nature of chemical reactions the laws of chemistry models and theories of chemistry chemical nomenclature mathematics and data analysis and instrumentation which are important ingredients of what one might call the nature of chemistry Using the history of the iron III thiocyanate reaction as a basis the book s main objective is to explore how chemistry develops its own knowledge base how it assesses the reliability of that base and how some important tools of the trade have been brought to bear on a chemical reaction to achieve understanding a worthwhile goal of any historical investigation **Geothermal Energy Update** ,1981-12

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