

ANALYSIS OF OBSERVATION(S)

Please use the example of an Analysis of Observation provided as a template for all analysis papers.

Step 1: Summarize what was observed.

1.A. Background Details:

Develop a chart that identifies:

- a. Site Visit Number(s)
- b. Day(s) of the week
- c. Date(s)
- d. Grade level(s)
- e. Subject(s)

1.B. Students

On average, how many students were present during the observation(s)? Based on the observation(s), describe the diversity of the students. Refer to IPTS 1.A for details regarding diversity.

1.C. Adults

How many other adults were in the room? Describe their roles (, parent, co-teacher, teacher's aide, student teacher, special education teacher).

1.D. Summary of Observation(s)

Using your Site Visit Notes and focusing on the appropriate Performance Activity/ies, provide 3–4 sentences summarizing what was observed in each Site Visit. Using the performance activities, provide clear examples which support the given IPTS.

Remember that a summary focuses on retelling the main events of the observation in your own words. No opinion or analysis should be part of your summary.

Step 2: Analyze observation(s) and the connection(s) to the standard/indicator. In this analysis, closely examine the observation(s) to understand the parts and ways they work together. Focus on the connection(s) between the observation(s) and the explicit standard/indicator the observation(s) addressed. Be sure to support all claims with specific examples from the observation(s).

Step 3: Analyze what was observed and its impact on student learning. Using classroom data and anecdotal data from the observation(s), **analyze** the way in which what was observed impacts student learning. Consider the way specific characteristics of students in the observed class(es) (, developmental characteristics, previously demonstrated academic performance/ability and special needs, socio-economic profile, ethnicity/cultural/gender make-up, special needs and language) impact decisions regarding instruction and assessments, especially in terms of the explicit standard on which the observation(s) focused. Provide specific numbers whenever possible. Be sure to support all claims with specific examples from the observation(s).

Observation Analysis Paper

KJ Lindholm-Leary



Observation Analysis Paper:

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a net ionic equation is the most accurate representation of the actual chemical process that occurs writing these equations requires a familiarity with solubility rules acid base reactivity

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apr 25 2017 this is understood when you consider the two chemicals appear unchanged before and after the reaction therefore the net ionic equation is $2\text{Cl}^- + 2\text{Ag}^+$

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ions that remain essentially unchanged during a reaction are called spectator ions this means you can ignore them when you write the ionic equation you only need to model how the solid

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1 K_3PO_4 aq + $\text{Al}(\text{NO}_3)_3$ aq $\rightarrow$ AlPO_4 s + 3KNO_3 aq
 3Cl^- aq + Al^{3+} aq $\rightarrow$ AlCl_3 s
 3K^+ aq + 3PO_4^{3-} aq + 3Al^{3+} aq + 9NO_3^- aq $\rightarrow$ $3\text{K}_3\text{PO}_4$ s + $9\text{Al}(\text{NO}_3)_3$ aq
 net ionic: Al^{3+} aq + 3PO_4^{3-} aq $\rightarrow$ AlPO_4 s
 aq 3 mg s

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the first step to writing a net ionic equation is balancing the chemical equation present let s use the reaction between sodium chloride and silver nitrate as an example next we write the

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highlights learning objectives by the end of this section you will be able to derive chemical equations from narrative descriptions of chemical reactions write and balance chemical

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the net ionic equation is the chemical equation that shows only those elements compounds and ions that are directly involved in the chemical reaction notice that in writing the net ionic

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what remains when the spectator ions are removed is called the net ionic equation which represents the actual chemical change occurring between the ionic compounds Cl^-aq

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this net ionic equation indicates that solid silver chloride may be produced from dissolved chloride and silver I^+ ions regardless of the source of these ions these molecular and

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write the overall chemical equation the complete ionic equation and the net ionic equation for the reaction of aqueous silver fluoride with aqueous sodium phosphate to give solid silver

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the complete ionic equation for this reaction is as follows equation 4 $13\text{Ag}^+\text{aq} + 2\text{F}^-\text{aq} + 2\text{NH}_4^+\text{aq} + \text{Cr}_2\text{O}_7^{2-}\text{aq} \rightarrow \text{Ag}_2\text{Cr}_2\text{O}_7\text{s} + 2\text{NH}_4^+\text{aq} + 2\text{F}^-\text{aq}$ because two NH_4^+aq

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