

Model 2 – Ternary Acids (Oxyacids)

Acid	Name of Acid in Aqueous Solution	Cation (+)	Polyatomic Anion (-)	Polyatomic Anion Name
HClO ₃	Chloric acid	H ₃ O ⁺	ClO ₃ ⁻¹	Chlorate
H ₂ SO ₃	Sulfurous acid	2 H ₃ O ⁺	SO ₃ ⁻²	Sulfite
H ₂ SO ₄	Sulfuric acid	2 H ₃ O ⁺	SO ₄ ⁻²	Sulfate
H ₃ PO ₃	Phosphorous acid	3 H ₃ O ⁺	PO ₃ ⁻³	Phosphite
H ₃ PO ₄	Phosphoric acid	3 H ₃ O ⁺	PO ₄ ⁻³	Phosphate
HNO ₃	Nitric acid	H ₃ O ⁺	NO ₃ ⁻¹	Nitrate
HNO ₂	Nitrous acid	H ₃ O ⁺	NO ₂ ⁻¹	Nitrite
H ₂ CO ₃	Carbonic acid	2 H ₃ O ⁺	CO ₃ ⁻²	Carbonate

6. Look at the formulas of the ternary acids in Model 2.

a. How are ternary acids different from binary acids in their structure?

Ternary acids contain three elements and they also contain the element oxygen.

b. What number do you think the prefix "ter-" refers to?

"ter" refers to three

7. When ternary acids are mixed with water, ions will form. Fill in the table above with the formulas and names of the anions.

8. Examine the pairs of ternary acids in Model 2 that contain sulfur, phosphorus, and nitrogen. Each pair has one acid that ends in "-ic" and another that ends in "-ous." These endings are related to the name of the polyatomic anion found in the acid ("-ate" or "-ite"). Complete the statements below with the correct acid name ending.

Polyatomic anion ending is "-ate" → acid name ending is "ic"

Polyatomic anion ending is "-ite" → acid name ending is "ous"

9. If the prefix "hydro-" were used to name a ternary acid, what problem would this create when naming HClO₃?

HClO₃⁻¹ has the name of chlorate. "ate" must be replaced with "ic". If "hydro" is used as a prefix, it would be hydrochloric acid. This has the formula HCl, not HClO₃.

10. Write a rule for naming ternary acids.

- The prefix "hydro" is not used in the name when an acid contains more than two elements and one of the elements is oxygen.

- "ate" is replaced with "ic" and "ite" is replaced with "ous".

11. Predict the formula for chlorous acid.



12. Circle the acid(s) below that would be named beginning with the prefix "hydro-."



Polyatomic Ions Pogil Worksheet Answers

Aharon Loewenschuss, Y. Marcus



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Polyatomic Ions for Surface Analysis and Modification Erick Ryan Fuoco, 2003 *Investigations Into the Origins of Polyatomic Ions in Inductively Coupled Plasma-mass Spectrometry* Sally M. McIntyre, 2010 *The Particulate Nature of Polyatomic Ions*, 2007 **Fragmentation of Diatomic and Polyatomic Ions in the Gas Phase** Christopher John Proctor, 1981 High Resolution Studies of the Origins of Polyatomic Ions in Inductively Coupled Plasma-Mass Spectrometry Jill Wisniewski Ferguson, 2006 The inductively coupled plasma ICP is an atmospheric pressure ionization source. Traditionally the plasma is sampled via a sampler cone. A supersonic jet develops behind the sampler and this region is pumped down to a pressure of approximately one Torr. A skimmer cone is located inside this zone of silence to transmit ions into the mass spectrometer. The position of the sampler and skimmer cones relative to the initial radiation and normal analytical zones of the plasma is key to optimizing the useful analytical signal. 1 The ICP both atomizes and ionizes the sample. Polyatomic ions form through ion molecule interactions either in the ICP or during ion extraction. 1 Common polyatomic ions that inhibit analysis include metal oxides. MO sup adducts with argon, the gas most commonly used to make up the plasma, and hydride species. While high resolution devices can separate many analytes from common interferences, this is done at great cost in ion transmission efficiency: a loss of 99% when using high versus low resolution on the same instrument. 2 Simple quadrupole devices, which make up the bulk of ICP MS instruments in existence, do not present this option. Therefore, if the source of polyatomic interferences can be determined and then manipulated, this could potentially improve the figures of merit on all ICP MS devices, not just the high resolution devices often utilized to study polyatomic interferences.

Investigations Into the Origins of Polyatomic Ions in Inductively Coupled Plasma-mass Spectrometry, 2010 An inductively coupled plasma mass spectrometer (ICP MS) is an elemental analytical instrument capable of determining nearly all elements in the periodic table at limits of detection in the parts per quadrillion and with a linear analytical range over 8-10 orders of magnitude. Three concentric quartz tubes make up the plasma torch. Argon gas is spiraled through the outer tube and generates the plasma, powered by a looped load coil operating at 27.1 or 40.6 MHz. The argon flow of the middle channel is used to keep the plasma above the innermost tube, through which solid or aqueous sample is carried in a third argon stream. A sample is progressively desolvated, atomized, and ionized. The torch is operated at atmospheric pressure. To reach the reduced pressures of mass spectrometers, ions are extracted through a series of two approximately one millimeter wide circular apertures set in water-cooled metal cones. The space between the cones is evacuated to approximately one torr. The space behind the second cone is pumped down to or near to the pressure needed for the mass spectrometer. MS. The first cone, called the sampler, is placed directly in the plasma plume and its position is adjusted to the point where atomic ions are most abundant. The hot plasma gas expands through the sampler orifice and in this expansion is placed the second cone, called the skimmer. After the skimmer, traditional MS designs are employed, i.e., quadrupoles, magnetic sectors, time of flight.

ICP MS is the leading trace element analysis technique. One of its weaknesses are polyatomic ions. This dissertation has added to the fundamental understanding of some of these polyatomic ions, their origins and behavior. Although mainly continuing the work of others, certain novel approaches have been introduced here. Chapter 2 includes the first reported efforts to include high temperature corrections to the partition functions of the polyatomic ions in ICP MS. This and other objections to preceding papers in this area were addressed. Errors in the measured T_{gas} values were found for given errors in the experimental and spectroscopic values. The ionization energy of the neutral polyatomic ion was included in calculations to prove the validity of ignoring more complicated equilibria. Work was begun on the question of agreement between kinetics of the plasma and interface and the increase and depletion seen in certain polyatomic ions. This dissertation was also the first to report day to day ranges for T_{gas} values and to use a statistical test to compare different operating conditions. This will help guide comparisons of previous and future work. Chapter 4 was the first attempt to include the excited electronic state 2 in the partition function of ArO as well as the first to address the different dissociation products of the ground and first electronic levels of ArO. Chapter 5 reports an interesting source of memory in ICP MS that could affect mathematical corrections for polyatomic ions. For future work on these topics, I suggest the following experiments and investigations. Clearly not an extensive list, they are instead the first topics curiosity brings to mind.

1. Measurement of T_{gas} values when using the flow injection technique of Appendix B. It was believed that there was a fundamental difference in the plasma when the auto sampler was used versus a continuous injection. Is this reflected in T_{gas} values?
2. The work of Chapter 3 can be expanded and supplemented with more trials, new cone materials, i.e. copper, stainless steel, and more cone geometries. Some of this equipment is already present in the laboratory; others could be purchased or made.
3. T_{gas} values from Chapter 3 could be correlated with instrument pressures during the experiment. Pressures after the skimmer cone were recorded for many days but have yet to be collated with the measured T_{gas} values.
4. *Fragmentation of Diatomic and Polyatomic Ions in the Gas Phase*, Christopher John Proctor, 1981
- Standard Thermodynamic Functions of Gaseous Polyatomic Ions at 100-1000 K*, Aharon Loewenschuss, Y. Marcus, 1987
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