Name

Date ___/___Period __

Chapter 5 • Periodicity Activity

ATOMIC RADII

Directions: Using the given data for the sizes of atoms (atomic radii) below,

- 1. Create a line graph (use a ruler) by graphing Period 2's data on the graph below.
- In a different color, create a line graph Period 3's data on the same graph.
- Label each axis, and include a key/legend.
- 4. Answer all questions.

		2	1.3	1.4	1.5	16	17	1.8	
I	30	Atomic Radii							
I	Li 123	Ble 89	B 88	C 77	N 70	O 66	F 64	Ne	
	Na 157	Mg 1.3-6	AI 125	Si 117	P 110	S 104	C1 99	Ar	
1	- K 2003	Ca. 174	Ga 125	Gie 122	As 121	Se 117	Be 114	Kr	
Ī	Rb 216	Sr 192	In 150	Sn 140	Sb 141	Te: 13-7	133	Xe	
	Cs 235	Ba 198	T1 155	Pb 154	Bi 152	Po 153	At	Ran	

Note: Mark the vertical axis from 0 - 160 (by 20's)

Trend in Atomic Size Across the Periodic Table

						-
	A STATE OF THE STA		200	The second secon		
Li	Be Mg	18	C	24	8	.16
Na	9.00	8.0	56	100	40	4000
- 276 CH	200 miles	2764	256	100	79	9-3

Horizontal Trend Questions

1. What is the trend in atomic size as you go across the periodic table?

2. This is true because there are more ______ (protons / electrons / layers of electrons) pulling the electron cloud in toward the nucleus.

Vertical Trend Questions

3. Examine the sizes of the Group 1 elements (H, Li, Na, K, Rb, Cs). What is the trend in atomic size as you go down a column of the periodic table?

4. This is true because there are more ______ (protons / electrons / layers of electrons) making the electron cloud larger.

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DP Hallahan

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