Quantum Number Problems

Use the rules for quantum numbers and notes to answer the questions below:

Rules: (1) n = 1, 2, ... 0

(3)
$$m_i = -l$$
, $-l+1$, ...0,+ l (e.g., when $l = 1$, $m_i = -1$, 0 , $+1$)

- (4) The number of nodal planes (Np) is equal to the angular momentum quantum number. This means /= Np.
- (5) The total number of nodes is $N_T = N_p + N_s$, where N_s is the number of spherical nodes (radial nodes). $N_T = n-1$ where n is the principal quantum number. So $n = N_p + N_s + 1$

From these orbitals: 2s 3d 4p 5f 6d 7p

Fill in the blanks with the correct answer(s):

(answers may be used more than once or not at all; none is a possible answer)

The orbital(s) with n = 6

____ 4. The

The orbital(s) with 0 planar nodes

2. The orbital(s) with /= 1

5. The orbital(s) with 3 spherical nodes

The orbital(s) that can have m_i = -2.

6. The orbital(s) with a total of 5 nodes

Problems On Quantum Numbers

Roman Wölfel

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