

## List of Common Polyatomic Ions

### Cations:

| <u>Symbol</u>   | <u>Name</u> | <u>Charge</u> | <u>Symbol</u>      | <u>Name</u>   | <u>Charge</u> |
|-----------------|-------------|---------------|--------------------|---------------|---------------|
| $\text{NH}_4^+$ | Ammonium    | +1            | $\text{Hg}_2^{+2}$ | Dimercury (I) | +2            |

### Anions:

| <u>Symbol</u>                      | <u>Name</u>  | <u>Charge</u> | <u>Symbol</u>                | <u>Name</u> | <u>Charge</u> |
|------------------------------------|--------------|---------------|------------------------------|-------------|---------------|
| $\text{C}_2\text{H}_3\text{O}_2^-$ | Acetate      | -1            | $\text{CO}_3^{-2}$           | Carbonate   | -2            |
| $\text{HCO}_3^-$                   | Bicarbonate  | -1            | $\text{CrO}_4^{-2}$          | Chromate    | -2            |
| $\text{BrO}_3^-$                   | Bromate      | -1            | $\text{Cr}_2\text{O}_7^{-2}$ | Dichromate  | -2            |
| $\text{BrO}_2^-$                   | Bromite      | -1            | $\text{C}_2\text{O}_4^{-2}$  | Oxalate     | -2            |
| $\text{ClO}_4^-$                   | Perchlorate  | -1            | $\text{O}_2^{-2}$            | Peroxide    | -2            |
| $\text{ClO}_3^-$                   | Chlorate     | -1            | $\text{SO}_4^{-2}$           | Sulfate     | -2            |
| $\text{ClO}_2^-$                   | Chlorite     | -1            | $\text{SO}_3^{-2}$           | Sulfite     | -2            |
| $\text{ClO}^-$                     | Hypochlorite | -1            | $\text{S}_2\text{O}_3^{-2}$  | Thiosulfate | -2            |
| $\text{CN}^-$                      | Cyanide      | -1            | $\text{SiO}_3^{-2}$          | Silicate    | -2            |
| $\text{SCN}^-$                     | Thiocyanate  | -1            | $\text{MoO}_4^{-2}$          | Molybdate   | -2            |
| $\text{OH}^-$                      | Hydroxide    | -1            | $\text{WO}_4^{-2}$           | Tungstate   | -2            |
| $\text{NO}_3^-$                    | Nitrate      | -1            | $\text{BO}_3^{-3}$           | Borate      | -3            |
| $\text{NO}_2^-$                    | Nitrite      | -1            | $\text{PO}_4^{-3}$           | Phosphate   | -3            |
| $\text{MnO}_4^-$                   | Permanganate | -1            | $\text{PO}_3^{-3}$           | Phosphite   | -3            |
| $\text{HSO}_4^-$                   | Bisulfate    | -1            | $\text{AsO}_4^{-3}$          | Arsenate    | -3            |
| $\text{HSO}_3^-$                   | Bisulfite    | -1            | $\text{AsO}_3^{-3}$          | Arsenite    | -3            |
| $\text{IO}_3^-$                    | Iodate       | -1            |                              |             |               |
| $\text{IO}_4^-$                    | Periodate    | -1            |                              |             |               |
| $\text{C}_6\text{H}_5\text{COO}^-$ | Benzoate     | -1            |                              |             |               |

| <u>Symbol</u>                 | <u>Name</u>            | <u>Charge</u> |
|-------------------------------|------------------------|---------------|
| $\text{H}_2\text{PO}_4^{-1}$  | Dihydrogen phosphate   | -1            |
| $\text{HPO}_4^{-2}$           | Hydrogen phosphate     | -2            |
| $\text{Fe}(\text{CN})_6^{-3}$ | Hexacyanoferrate (III) | -3            |
| $\text{Fe}(\text{CN})_6^{-4}$ | Hexacyanoferrate (II)  | -4            |

# Polyatomic Ions Packet Answers

**Jill Wisnewski Ferguson**



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