Name ___Key___
Period ___ Date ___/___/___

7&8 • Atomic Structure & Periodicity

PRACTICE TEST

Rhc = A (handwritten)

> $A = 2.18 \times 10^{-18}$ J $h = 6.626 \times 10^{-34}$ J·s
>
> $R = 1.097 \times 10^{-7}$ m^{-1} $c = 3.00 \times 10^{8}$ m·s^{-1}
>
> mass of an electron $= 9.11 \times 10^{-31}$ kg

1. What wavelength corresponds to a frequency of 8.22×10^{9} Hz? $\lambda = \frac{c}{v}$
   a) 0.307 m   d) 0.110 m
   b) 0.0365 m $\frac{3.00 \times 10^{8}}{8.22 \times 10^{9}\ s^{-1}}$ e) 27.4 m
   c) 0.122 m $= .036496...$

   (b circled)

2. A radio station transmits at 110 MHz (110×10^{6} Hz). What wavelength is this radio wave?
   a) 3.65×10^{-5} m   c) 3.81×10^{-5} m
   b) 3.30 m   d) 2.73 m
   $\lambda = \frac{c}{v} = \frac{3.00 \times 10^{8}\ m \cdot s^{-1}}{110 \times 10^{6}\ s^{-1}} = 2.73\ m$

   (d circled)

3. Which one of the following is NOT a proper unit for frequency?
   a) Hz
   b) s^{-1}
   c) m·s^{-1} velocity
   d) $\frac{1}{sec}$

   (c circled)

4. Calculate the wavelength of the fourth line in the Balmer series (the visible series) of the hydrogen spectrum. $n=6 \rightarrow n=2$
   $E_6 = \frac{-2.18 \times 10^{-18} J}{36} - E_2 = \frac{-2.18 \times 10^{-18} J}{4}$
   a) 0.12334 m   d) 4.1029×10^{-7} m
   b) 24.373 m   e) 36.559 m
   c) 2.7353×10^{-7} m $E = \frac{hc}{\lambda}$ $\lambda = \frac{hc}{E}$
   $\Delta E = 4.84 \times 10^{-19} J$ $= 4.1 \times 10^{-7}$ m

   (d circled)

5. What is the relationship between the energy of a photon of light and its frequency?
   a) $E = v$ d) $E = \frac{1}{hv}$
   b) $E = \frac{h}{v}$ e) $E = \frac{v}{h}$
   c) $E = hv$

   (c circled)

6. What is the energy needed to raise an electron in the hydrogen atom from the second energy level to the third energy level? $\Delta E = E_3 - E_2$
   a) 1.52×10^{4} J   d) 4.48×10^{-19} J
   b) 3.63×10^{-19} J   e) 3.03×10^{-19} J
   c) 2.18×10^{-19} J
   $\frac{-2.18 \times 10^{-18} J}{9} - \frac{-2.18 \times 10^{-18} J}{4}$

   (e circled)

7. What is the de Broglie wavelength of an electron moving at 80.0% the speed of light.
   a) 3.03×10^{-12} m   c) 3.30×10^{11} m
   b) 2.42×10^{-12} m   d) 1.59×10^{-25} m
   $\lambda = \frac{h}{mv} = \frac{6.626 \times 10^{-34} J \cdot s}{9.11 \times 10^{-31} kg \times .8 \times 3.00 \times 10^{8}}$

   (a circled)

8. What resultant is expected from the interference of the two waves shown below?

   a)
   b)
   c)
   d)

   (c circled)

9. Which quantum number determines the subshell occupied by an electron (s, p, d, f, etc.)?
   a) n c) m_ℓ
   b) ℓ d) m_s

   (b circled)

Problem 23a Quantum Energy Answers

Mark Steyvers

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