

**Gas Laws Worksheet**

$$\text{atm} = 760.0 \text{ mm Hg} = 101.3 \text{ kPa} = 760.0 \text{ torr}$$

**Boyle's Law Problems:**

$$P_1 V_1 = P_2 V_2$$

1. If 22.5 L of nitrogen at 748 mm Hg are compressed to 725 mm Hg at constant temperature. What is the new volume?

$$(748 \text{ mmHg})(22.5 \text{ L}) = (725 \text{ mmHg}) V_2$$

$$V_2 = \frac{(748 \text{ mmHg})(22.5 \text{ L})}{(725 \text{ mmHg})}$$

$$V_2 = 23.2 \text{ L}$$

2. A gas with a volume of 4.0 L at a pressure of 205 kPa is allowed to expand to a volume of 12.0 L. What is the pressure in the container if the temperature remains constant?

$$(4.0 \text{ L})(205 \text{ kPa}) = (12.0 \text{ L}) P_2$$

$$P_2 = \frac{(4.0 \text{ L})(205 \text{ kPa})}{12.0 \text{ L}}$$

$$P_2 = 68.3 \text{ kPa}$$

3. What pressure is required to compress 196.0 liters of air at 1.00 atmosphere into a cylinder whose volume is 26.0 liters?

$$(196.0 \text{ L})(1.00 \text{ atm}) = (26.0 \text{ L}) P_2$$

$$P_2 = \frac{(196.0 \text{ L})(1.00 \text{ atm})}{26.0 \text{ L}}$$

$$P_2 = 7.54 \text{ atm}$$

4. A 40.0 L tank of ammonia has a pressure of 12.7 kPa. Calculate the volume of the ammonia if its pressure is changed to 8.4 kPa while its temperature remains constant.

$$(40.0 \text{ L})(12.7 \text{ kPa}) = (8.4 \text{ kPa}) V_2$$

$$V_2 = \frac{(40.0 \text{ L})(12.7 \text{ kPa})}{8.4 \text{ kPa}}$$

$$V_2 = 60.5 \text{ L}$$

# Practice With Gas Law Calculations Iii

**Dean R. Hess, Neil R.  
MacIntyre, William F. Galvin**



### **Practice With Gas Law Calculations Iii:**

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## **Table of Contents Practice With Gas Law Calculations Iii**

1. Understanding the eBook Practice With Gas Law Calculations Iii
  - The Rise of Digital Reading Practice With Gas Law Calculations Iii
  - Advantages of eBooks Over Traditional Books
2. Identifying Practice With Gas Law Calculations Iii
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Practice With Gas Law Calculations Iii
  - User-Friendly Interface
4. Exploring eBook Recommendations from Practice With Gas Law Calculations Iii
  - Personalized Recommendations
  - Practice With Gas Law Calculations Iii User Reviews and Ratings
  - Practice With Gas Law Calculations Iii and Bestseller Lists
5. Accessing Practice With Gas Law Calculations Iii Free and Paid eBooks
  - Practice With Gas Law Calculations Iii Public Domain eBooks
  - Practice With Gas Law Calculations Iii eBook Subscription Services
  - Practice With Gas Law Calculations Iii Budget-Friendly Options
6. Navigating Practice With Gas Law Calculations Iii eBook Formats

- ePub, PDF, MOBI, and More
- Practice With Gas Law Calculations Iii Compatibility with Devices
- Practice With Gas Law Calculations Iii Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Practice With Gas Law Calculations Iii
  - Highlighting and Note-Taking Practice With Gas Law Calculations Iii
  - Interactive Elements Practice With Gas Law Calculations Iii
- 8. Staying Engaged with Practice With Gas Law Calculations Iii
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Practice With Gas Law Calculations Iii
- 9. Balancing eBooks and Physical Books Practice With Gas Law Calculations Iii
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Practice With Gas Law Calculations Iii
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Practice With Gas Law Calculations Iii
  - Setting Reading Goals Practice With Gas Law Calculations Iii
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Practice With Gas Law Calculations Iii
  - Fact-Checking eBook Content of Practice With Gas Law Calculations Iii
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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