

## Specific Heat Capacity Practice Problems

Name: \_\_\_\_\_

1. When 3.0 kg of water is cooled from 80.0°C to 10.0°C, how much heat energy is lost?

$$c_{p,w} = Q / m \cdot \Delta T$$

$$Q = c_{p,w} (m) \Delta T$$

$$= (4186)(3)(10 - 80)$$

$$\boxed{-879,060 \text{ J}}$$

2. How much heat is needed to raise a 0.30 kg piece of aluminum from 30.°C to 150°C?

$$Q = c_{p,a} (m_a) (\Delta T_a)$$

$$= 899 (.3) (150 - 30)$$

$$\boxed{+32,364 \text{ J}}$$

3. Calculate the temperature change when:

- a) 10.0 kg of water loses 232 kJ of heat.

$$232 \text{ kJ} \cdot \frac{1000 \text{ J}}{1 \text{ kJ}} = 232,000 \text{ J}$$

$$-232,000 \text{ J} = (4186)(10.0 \text{ kg}) \Delta T$$

$$\frac{-232,000}{4186} = \Delta T$$

$$\Delta T = -55.4^\circ\text{C}$$

- b) 1.96 kJ of heat are added to 500. g of copper.

$$1.96 \text{ kJ} \cdot \frac{1000 \text{ J}}{1 \text{ kJ}} = 1960 \text{ J}$$

$$500 \text{ g} \cdot \frac{1 \text{ kg}}{1000 \text{ g}} = .5 \text{ kg}$$

$$1960 \text{ J} = (387)(.5 \text{ kg}) \Delta T$$

$$\frac{1960}{193.5} = \Delta T$$

$$\Delta T = 10.1^\circ\text{C}$$

4.  $2.52 \times 10^4 \text{ J}$  of heat are added to 2.0 kg of mercury to reach a final temperature of 130°C. What was the initial temperature of the mercury?

$$T_f = 130^\circ\text{C}$$

$$Q = 2.52 \times 10^4$$

$$m = 2.0 \text{ kg}$$

$$c_{p,m} = 138$$

$$2.52 \times 10^4 = (138)(2)(130 - T_i)$$

$$2.52 \times 10^4 = 276(130 - T_i)$$

$$91.3 = 130 - T_i$$

$$-38.7 = -T_i$$

$$\boxed{38.7^\circ\text{C} = T_i}$$

# Practice Specific Heat Problems With Answers

**LM Reder**



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subject making it easier for students to solve for the exams      *Study Guide with Answers to Selected Problems* H. Stephen Stoker, Danny V. White, 2003-04 This supplement includes for each chapter a brief overview activities and practice problems to reinforce skills and a practice test The answers section includes answers for all odd numbered end of chapter exercises

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In an electronic era where connections and knowledge reign supreme, the enchanting power of language has are more apparent than ever. Its power to stir emotions, provoke thought, and instigate transformation is really remarkable. This extraordinary book, aptly titled "**Practice Specific Heat Problems With Answers**," published by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound impact on our existence. Throughout this critique, we shall delve to the book is central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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