

Chemistry 1220
Practice Exam 3
7 July, 2003

Instructions:

Do not begin until 8:40 AM.

This exam should have 25 questions. Each question is worth 4 points for a total of 100 points. A periodic table should be on the reverse side of this page.

Use only a #2 pencil to mark the best answer on the scantron Sheet. You will turn in only the scantron sheet and your homework set 2. You should keep the rest of the exam.

On the scantron sheet, fill out your name and SS# and be sure to fill out the corresponding bubbles.

Information:

$$K_w = [\text{H}_3\text{O}^+][\text{OH}^-] = 1.0 \times 10^{-14} \text{ (at } 25^\circ\text{C)}$$

$$K_a \times K_b = K_w$$

$$\text{pH} = \text{p}K_a + \log ([\text{A}^-]/[\text{HA}])$$

$$\text{pH} + \text{pOH} = 14$$

$$\Delta S_{\text{sys}} = q_{\text{rev}}/T \text{ (at constant } T)$$

$$\Delta G = \Delta H - T\Delta S$$

$$\Delta G = \Delta G^\circ + RT \ln Q$$

$$\Delta G^\circ = -RT \ln K_{\text{eq}}$$

$$1\text{V} = 1\text{J/C}$$

$$1\text{W} = 1\text{J/s}$$

$$\Delta G = -nFE$$

$$\Delta G^\circ = -nFE^\circ$$

$$1\text{ F} = 96500\text{ C/mol} = 96,500\text{ J/(V} \times \text{mol)}$$

$$E^\circ_{\text{cell}} = E^\circ_{\text{red(cathode)}} - E^\circ_{\text{red(anode)}}$$

$$E^\circ_{\text{cell}} = E^\circ_{\text{red(cathode)}} + E^\circ_{\text{ox(anode)}}$$

$$E = E^\circ - (0.0592/n) \log Q$$

$$\log K_{\text{eq}} = (nE^\circ)/0.0592$$

$$w = nFE_{\text{ext}}$$

$$R = 8.314\text{ J/(mol} \times \text{K)}$$

$$K = ^\circ\text{C} + 273.15$$

$$\Delta H^\circ = \sum n H^\circ(\text{products}) - \sum m H^\circ(\text{reactants})$$

$$\Delta S^\circ = \sum n S^\circ(\text{products}) - \sum m S^\circ(\text{reactants})$$

$$\Delta G^\circ = \sum n \Delta G^\circ_f(\text{products}) - \sum m \Delta G^\circ_f(\text{reactants})$$

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JA Banks



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