

## Density Practice Problems Worksheet #2

### Answer Key

1. You have a rock with a volume of  $15\text{cm}^3$  and a mass of 45 g. What is its density?

$$\text{Density} = m/v$$

$$\text{Density} = 45\text{g}/15\text{cm}^3$$

$$\text{Density} = 3.0\text{g}/\text{cm}^3$$

2. You have a different rock with a volume of  $30\text{cm}^3$  and a mass of 60g. What is its density?

$$\text{Density} = m/v$$

$$\text{Density} = 60\text{g}/30\text{cm}^3$$

$$\text{Density} = 2.0\text{g}/\text{cm}^3$$

3. In the above two examples which rock is more dense? Why?

Rock sample #1, because it has a density of  $3.0\text{g}/\text{cm}^3$ , which is greater than rock sample #2's density of  $2.0\text{g}/\text{cm}^3$

4. Calculate the mass of a liquid with a density of  $3.2\text{ g/mL}$  and a volume of 25 mL.

$$\text{Mass} = D \times V$$

$$\text{Mass} = 3.2\text{g/mL} \times 25\text{ mL}$$

$$\text{Mass} = 80.0\text{ g}$$

5. Calculate the density of a 500g rectangular block with the following dimensions: length=8 cm, width=6 cm, height=5 cm.

$$\text{Density} = m/v$$

$$\text{Density} = 500\text{g}/240\text{cm}^3 (8\text{cm} \times 6\text{cm} \times 5\text{cm})$$

$$\text{Density} = 2.083, \text{ rounded to nearest tenth} = 2.1\text{ g}/\text{cm}^3$$

6. Calculate the mass of a solid metal block with a density of  $2.6\text{ g}/\text{cm}^3$ , a width of 1.8 cm, a length of 4 cm, and a height of 2.3 cm.

$$\text{Mass} = D \times V$$

$$\text{Mass} = 2.6\text{g}/\text{cm}^3 \times (1.8\text{cm} \times 4\text{cm} \times 2.3\text{ cm})$$

$$\text{Mass} = 2.6\text{g}/\text{cm}^3 \times 16.56\text{cm}^3$$

$$\text{Mass} = 43.056\text{g} - \text{rounded to nearest tenth} = 43.1\text{ g}$$

# Practice Problems And Answers For Chemistry Density

**Donald J. Wink, Sharon Fetzer-  
Gislason, Sheila McNicholas**



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Khomtchouk,2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics The focus is not to recap or review the theoretical concepts well described in the available texts Instead the topics and descriptions in this book make available specific detailed step by step methods and procedures for solving the major types of problems in general chemistry Explanations instructional process sequences solved examples and completely solved practice problems are greatly expanded containing significantly more detail than can usually be devoted to in a comprehensive text Many chapters also provide alternative viewpoints as an aid to understanding Key Features The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester Each is written in a specific and detailed step by step process for problem solving whether mathematical or conceptual Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts Includes a chapter designed to eliminate confusion concerning acid base reactions which often persists through working with acid base equilibrium Many chapters provide alternative viewpoints as an aid to understanding This book addresses a very real need for a large number of incoming freshman in STEM fields

**The Practice of Chemistry Study Guide & Solutions Manual** Pamela Mills,Amina El-Ashmawy,2003-04-14 Designed to help students understand the material better and avoid common mistakes Also includes solutions and explanations to odd numbered exercises

**General Chemistry** Donald A. McQuarrie,Stanley Gill,2011-06-15 This Fourth Edition of McQuarrie s classic text offers a thorough revision and a quantum leap forward from the previous edition Taking an atoms first approach it promises to be another

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Numerous examples are presented in every chapter to aid students and provide helpful self learning tools The topics are arranged throughout the textbook in a traditional approach to the subject with the primary audience being undergraduate students and advanced high school students of chemistry

Basic Concepts of Chemistry Leo J. Malone, Theodore O. Dolter, 2011-12-27 The 9th edition of Malone s Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter WileyPLUS sold separately from text

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**Ebook: Chemistry** Julia Burdge, 2014-10-16 Chemistry Third Edition by Julia Burdge offers a clear writing style written with the students in mind Julia uses her background of teaching hundreds of general chemistry students per year and creates content to offer more detailed explanation on areas where she knows they have problems With outstanding art a consistent problem solving approach

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## **Table of Contents Practice Problems And Answers For Chemistry Density**

1. Understanding the eBook Practice Problems And Answers For Chemistry Density
  - The Rise of Digital Reading Practice Problems And Answers For Chemistry Density
  - Advantages of eBooks Over Traditional Books
2. Identifying Practice Problems And Answers For Chemistry Density
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Practice Problems And Answers For Chemistry Density
  - User-Friendly Interface
4. Exploring eBook Recommendations from Practice Problems And Answers For Chemistry Density
  - Personalized Recommendations
  - Practice Problems And Answers For Chemistry Density User Reviews and Ratings
  - Practice Problems And Answers For Chemistry Density and Bestseller Lists
5. Accessing Practice Problems And Answers For Chemistry Density Free and Paid eBooks



- Practice Problems And Answers For Chemistry Density Public Domain eBooks
  - Practice Problems And Answers For Chemistry Density eBook Subscription Services
  - Practice Problems And Answers For Chemistry Density Budget-Friendly Options
6. Navigating Practice Problems And Answers For Chemistry Density eBook Formats
    - ePub, PDF, MOBI, and More
    - Practice Problems And Answers For Chemistry Density Compatibility with Devices
    - Practice Problems And Answers For Chemistry Density Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Practice Problems And Answers For Chemistry Density
    - Highlighting and Note-Taking Practice Problems And Answers For Chemistry Density
    - Interactive Elements Practice Problems And Answers For Chemistry Density
  8. Staying Engaged with Practice Problems And Answers For Chemistry Density
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Practice Problems And Answers For Chemistry Density
  9. Balancing eBooks and Physical Books Practice Problems And Answers For Chemistry Density
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Practice Problems And Answers For Chemistry Density
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Practice Problems And Answers For Chemistry Density
    - Setting Reading Goals Practice Problems And Answers For Chemistry Density
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Practice Problems And Answers For Chemistry Density
    - Fact-Checking eBook Content of Practice Problems And Answers For Chemistry Density
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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