



## NEW JERSEY CENTER FOR TEACHING & LEARNING

### *Aqueous Equilibria: Part 1 - Buffers & Titrations*

PSI AP Chemistry

Name \_\_\_\_\_

#### *Multiple Choice questions*

##### **Common-Ion Effect**

1) The pH of a solution that contains 0.818 M acetic acid ( $K_a = 1.77 \times 10^{-5}$ ) and 0.172 M sodium acetate is \_\_\_\_\_.

- A) 4.077      B) 5.434      C) 6.571      D) 8.370      E) 9.922

2) Calculate the pH of a solution prepared by dissolving 0.37 mol of formic acid ( $\text{HCO}_2\text{H}$ ) and 0.23 mol of sodium formate ( $\text{NaCO}_2\text{H}$ ) in water sufficient to yield 1.00 L of solution. The  $K_a$  of formic acid is  $1.77 \times 10^{-4}$ .

- A) 2.1      B) 10.46      C) 3.55      D) 2.31      E) 3.95

3) Calculate the pH of a solution prepared by dissolving 0.25 mol of benzoic acid ( $\text{C}_6\text{H}_5\text{O}_2\text{H}$ ) and 0.15 of sodium benzoate ( $\text{NaC}_6\text{H}_5\text{O}_2$ ) in water sufficient to yield 1.00 L of solution. The  $K_a$  of benzoic acid is  $6.50 \times 10^{-5}$ .

- A) 4.41      B) 2.40      C) 3.97      D) 10.08      E) 4.19

4) Calculate the pH of a solution prepared by dissolving 0.15 mol of benzoic acid ( $\text{HBz}$ ) and 0.30 mol of sodium benzoate in water sufficient to yield 1.00L of solution. The  $K_a$  of benzoic acid is  $6.50 \times 10^{-5}$ .

- A) 2.52      B) 3.89      C) 4.49      D) 10.16      E) 4.20

5) Calculate the pH of a solution prepared by dissolving 0.750 mol of  $\text{NH}_3$  and 0.250 mol of  $\text{NH}_4\text{Cl}$  in water sufficient to yield 1.00 L of solution. The  $K_b$  of ammonia is  $1.77 \times 10^{-5}$ .

- A) 5.22      B) 4.27      C) 9.73      D) 8.78      E) 0.89

6) The pH of a solution prepared by dissolving 0.350 mol of solid methylamine hydrochloride ( $\text{CH}_3\text{NH}_3\text{Cl}$ ) in 1.00L of 1.10M methylamine ( $\text{CH}_3\text{NH}_2$ ) is \_\_\_\_\_. The  $K_b$  for methylamine is  $4.4 \times 10^{-4}$ . (Assume the final volume is 1.00 L.)

- A) 1.86      B) 2.66      C) 10.15      D) 11.14      E) 10.64

##### **Buffers**

7) What does a buffer do?

- a) Keeps the pH of a solution constant.

# Practice Multiple Choice For Chem Aqueous Equilibria

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