

## Model 2 – Ternary Acids (Oxyacids)

| Acid                    | Name of Acid in Aqueous Solution | Cation (+)              | Polyatomic Anion (-) | Polyatomic Anion Name |
|-------------------------|----------------------------------|-------------------------|----------------------|-----------------------|
| $\text{HClO}_3$         | Chloric acid                     | $\text{H}_3\text{O}^+$  | $\text{ClO}_3^{-1}$  | Chlorate              |
| $\text{H}_2\text{SO}_3$ | Sulfurous acid                   | $2\text{H}_3\text{O}^+$ | $\text{SO}_3^{-2}$   | Sulfite               |
| $\text{H}_2\text{SO}_4$ | Sulfuric acid                    | $2\text{H}_3\text{O}^+$ | $\text{SO}_4^{-2}$   | Sulfate               |
| $\text{H}_3\text{PO}_3$ | Phosphorous acid                 | $3\text{H}_3\text{O}^+$ | $\text{PO}_3^{-3}$   | Phosphite             |
| $\text{H}_3\text{PO}_4$ | Phosphoric acid                  | $3\text{H}_3\text{O}^+$ | $\text{PO}_4^{-3}$   | Phosphate             |
| $\text{HNO}_3$          | Nitric acid                      | $\text{H}_3\text{O}^+$  | $\text{NO}_3^{-1}$   | Nitrate               |
| $\text{HNO}_2$          | Nitrous acid                     | $\text{H}_3\text{O}^+$  | $\text{NO}_2^{-1}$   | Nitrite               |
| $\text{H}_2\text{CO}_3$ | Carbonic acid                    | $2\text{H}_3\text{O}^+$ | $\text{CO}_3^{-2}$   | Carbonate             |

6. Look at the formulas of the ternary acids in Model 2.

a. How are ternary acids different from binary acids in their structure?

*Ternary acids contain three elements and they also contain the element oxygen.*

b. What number do you think the prefix "ter-" refers to?

*"ter" refers to three*

7. When ternary acids are mixed with water, ions will form. Fill in the table above with the formulas and names of the anions.

8. Examine the pairs of ternary acids in Model 2 that contain sulfur, phosphorus, and nitrogen. Each pair has one acid that ends in "-ic" and another that ends in "-ous." These endings are related to the name of the polyatomic anion found in the acid ("-ate" or "-ite"). Complete the statements below with the correct acid name ending.

Polyatomic anion ending is "-ate" → acid name ending is "ic"

Polyatomic anion ending is "-ite" → acid name ending is "ous"

9. If the prefix "hydro-" were used to name a ternary acid, what problem would this create when naming  $\text{HClO}_3$ ?

*$\text{ClO}_3^{-1}$  has the name of Chlorate. "ate" must be replaced with "ic". If "hydro" is used as a prefix, it would be hydrochloric acid. This has the formula  $\text{HCl}$ , not  $\text{HClO}_3$ .*

10. Write a rule for naming ternary acids.

*- The prefix "hydro" is not used in the name when an acid contains more than two elements and one of the elements is oxygen.*

*- "ate" is replaced with "ic" and "ite" is replaced with "ous".*

11. Predict the formula for chlorous acid.



12. Circle the acid(s) below that would be named beginning with the prefix "hydro-."



# Polyatomic Ions Worksheet Pogil

**Wenbin Ji**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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