

## Chemistry Chapter 8 Test

### Multiple Choice

Identify the choice that best completes the statement or answers the question.

- \_\_\_\_ 1. What type of bond would you expect in a compound of oxygen and sodium?  
☒ a. ionic  
 b. metallic  
 c. covalent  
 d. polar covalent
- \_\_\_\_ 2. The symbols  $\delta^-$  and  $\delta^+$  indicate which of the following?  
 a. electronegativity  
 b. electrostatic force  
 c. charge  
☒ d. partial charge
- \_\_\_\_ 3. Which of the following represents a nonpolar covalent molecule?  
 a. NaCl  
☒ b. Br<sub>2</sub>  
 c. HCN  
 d. Au
- \_\_\_\_ 4. The measure of an atom's ability to attract electrons to itself in a chemical bond is called \_\_\_\_\_.  
 a. electron affinity  
 b. polarity  
☒ c. electronegativity  
 d. ionization energy
- \_\_\_\_ 5. The polarity of a bond depends on the difference in the atoms' \_\_\_\_\_.  
 a. electron affinities  
☒ b. electronegativities  
 c. atomic radii  
 d. ionization energies
- \_\_\_\_ 6. What type of bond would you expect to form between nitrogen and potassium?  
☒ a. ionic  
 b. covalent  
 c. polar covalent  
 d. metallic
- \_\_\_\_  7. High electronegativity differences are found in \_\_\_\_\_ bonds.  
 a. metallic  
☒ b. ionic      metal - nonmetal  
 c. covalent  
 d. Pauling
- \_\_\_\_  8. To achieve an octet when combining with potassium, would an atom of a Group 16 element be expected to lose, gain, or share electrons?  
 a. lose  
☒ b. gain  
 c. share  
☒ d. This cannot be determined from the given information. x
- \_\_\_\_ 9. Which of the following does not exist as a diatomic molecule?  
☒ a. argon  
 b. hydrogen  
 c. chlorine  
 d. oxygen  
 H   N   O   F   Cl   Br   I  
 Gas : H<sub>2</sub>   N<sub>2</sub>   O<sub>2</sub>   F<sub>2</sub>   Cl<sub>2</sub>   Br<sub>2</sub>   I<sub>2</sub>
- \_\_\_\_ 10. Which of these is not a covalent molecule?  
 a. F<sub>2</sub>  
 b. CO  
☒ c. LiF  
 d. NH<sub>3</sub>
- \_\_\_\_ 11. A dash in a Lewis structure represents \_\_\_\_\_.  
☒ a. one electron  
 b. two electrons  
 c. four electrons  
 d. unpaired electrons
- \_\_\_\_ 12. The central atom in a Lewis structure is usually the atom \_\_\_\_\_.

# Odysseyware Chemistry Test Answers

**DJ Losen**



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