

- 1.1 Find the rate of reaction defined as

$$r_{O_2} = \frac{\text{mol } O_2 \text{ used}}{\text{sec} \cdot \text{m}^3 \text{ of tank}}$$

Evaluate terms

$$\bar{t} = \frac{V}{\bar{v}} \quad \text{or} \quad V = \bar{t} \bar{v}$$

or (Volume of treatment tanks) $= \left(\frac{1}{3} \text{ day}\right) \left(32000 \frac{\text{m}^3}{\text{day}}\right) = 10667 \text{ m}^3$

O_2 used $\left(200 \frac{\text{mg}}{\text{lit}}\right) \left(\frac{1 \text{ gm}}{1000 \text{ mg}}\right) \left(\frac{\text{mol}}{32 \text{ gm}}\right) \left(\frac{1000 \text{ lit}}{\text{m}^3}\right) \left(\frac{32000 \text{ m}^3}{\text{day}}\right) = 2 \times 10^5 \frac{\text{mol } O_2}{\text{day}}$

Thus the rate of reaction

$$\left. \begin{aligned} \frac{2.0 \times 10^5 \text{ mol } O_2 / \text{day}}{10667 \text{ m}^3} &= 18.75 \text{ mol/m}^3 \cdot \text{day} \\ &= 2.17 \times 10^{-4} \text{ mol/m}^3 \cdot \text{s} \end{aligned} \right\} \leftarrow$$

- 1.3 Find $-r'_{C_{20}H_{42}}$ and $-r^M_{C_{20}H_{42}}$... evaluate terms

$$V_{\text{cat}} = \frac{50000 \text{ kg}}{800 \text{ kg/m}^3} = 62.5 \text{ m}^3 \text{ of catalyst}$$

$$W_{\text{cat}} = 50000 \text{ kg}$$

$$\text{MW}_{C_{20}H_{42}} = [20(12) + 42(1)] \frac{1}{1000} = 0.282 \text{ kg/mol}$$

$$F_{\text{feed}} = (6000 \text{ m}^3/\text{day})(900 \text{ kg/m}^3) = 5400000 \text{ kg/day}$$

So $-\frac{dN_{C_{20}H_{42}}}{dt} = \left(\frac{5400000 \text{ kg/day}}{0.282 \text{ kg/day}}\right) \left(\frac{\text{day}}{24(3600) \text{ s}}\right) (0.6) = 133 \text{ mol reacted/s}$

Thus the rate of disappearance of $C_{20}H_{42}$

$$\left. \begin{aligned} -r' &= \frac{1}{W_{\text{cat}}} \frac{dN}{dt} = \frac{1}{50000} (133) = 0.0027 \text{ mol/kg cat} \cdot \text{s} \\ -r^M &= \frac{1}{V_{\text{cat}}} \frac{dN}{dt} = \frac{1}{62.5} (133) = 2.13 \text{ mol/m}^3 \text{ cat} \cdot \text{s} \end{aligned} \right\} \leftarrow$$

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Table of Contents Octave Levenspiel Solution Manual

1. Understanding the eBook Octave Levenspiel Solution Manual
 - The Rise of Digital Reading Octave Levenspiel Solution Manual
 - Advantages of eBooks Over Traditional Books
2. Identifying Octave Levenspiel Solution Manual
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Octave Levenspiel Solution Manual
 - User-Friendly Interface
4. Exploring eBook Recommendations from Octave Levenspiel Solution Manual
 - Personalized Recommendations

- Octave Levenspiel Solution Manual User Reviews and Ratings
- Octave Levenspiel Solution Manual and Bestseller Lists
- 5. Accessing Octave Levenspiel Solution Manual Free and Paid eBooks
 - Octave Levenspiel Solution Manual Public Domain eBooks
 - Octave Levenspiel Solution Manual eBook Subscription Services
 - Octave Levenspiel Solution Manual Budget-Friendly Options
- 6. Navigating Octave Levenspiel Solution Manual eBook Formats
 - ePub, PDF, MOBI, and More
 - Octave Levenspiel Solution Manual Compatibility with Devices
 - Octave Levenspiel Solution Manual Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Octave Levenspiel Solution Manual
 - Highlighting and Note-Taking Octave Levenspiel Solution Manual
 - Interactive Elements Octave Levenspiel Solution Manual
- 8. Staying Engaged with Octave Levenspiel Solution Manual
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Octave Levenspiel Solution Manual
- 9. Balancing eBooks and Physical Books Octave Levenspiel Solution Manual
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Octave Levenspiel Solution Manual
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Octave Levenspiel Solution Manual
 - Setting Reading Goals Octave Levenspiel Solution Manual
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Octave Levenspiel Solution Manual
 - Fact-Checking eBook Content of Octave Levenspiel Solution Manual

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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