Chemistry Chapter 8 Test

(b) two electrons

____ 12. The central atom in a Lewis structure is usually the atom ______

(denti)	fyr chw	cho	ice that best completes the statement or a	BUETHOOK.	s the question.
	1.	Wit	at type of bond would you expect in a co-	mp-our	nd of oxygen and sodium?
		(30)	ionic	C.	covalent
		Б.	metallic	d.	polar covalent
	2.	The	e symbols 5- and 5+ indicate which of the	follo	wing?
		4.	electronegativity	C.	change
		b.	electrostatic force	(d)	partial charge
-	3.	WE	with of the following represents a nonpola	r cowa	lent molecule?
		4.	NaCl	C.	HCN
		(E)	Br ₂	d.	Au
	43	The	e measure of an atom's ability to attract el	ectron	is to itself in a chemical bond is called
		a.	electron affinity	(6)	electronegativity
		b.	polarity	d.	ionization energy
	5.	The	e polarity of a bond depends on the different	moe in	the atoms'
		4.	electron affinities	· Cu	atomic radii
		(E)	electronegativities	d.	ionization energies
	1940			120000	N E
	663	and the same of th	sat type of bond would you expect to form		
		(3)	ionic		polar covalent
		ь.	covalent	d.	metallic
-	A.		th electronegativity differences are found		bonds.
		a.	noetallic		covalent
		(P)	ionic metal - narmetal	d.	Pauling C
	78	To	achieve an octet when combining with po-	s Cassini	m, would an atom of a Group 16 element be expected
		to I	ose, gain, or share electrons?	100	
		0	lose	K	505-0223
		(6)	gain		€ 13
		C.	shareX		
		(d)	This cannot be determined from the give	n imfo	rmation, k
_	9.	Wit	nich of the following does not exist as a di	ac ono is	c molecule? W N D F Cl Br I
		(P)	argon	· C.	chlorine
		Б.	hydrogen	d.	oxygen Gas M, N, O, F, Cl, S, I,
	10.	WE	with of these is not a covalent molecule?		
		46.	F2 0	(3)	Lif
		b.	CO+	d.	NH,
	11.	A. 0	lash in a Lewis structure represents		-
		20	one electron	100	four electrons

d. unpaired electrons

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Arshad Iqbal

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