

CHEMISTRY UNIT 2 REVISION

Energetics:

→ How to calculate general bond enthalpy?

- Bonds broken – bonds formed = H

→ What are the products of combustion?

- Carbon dioxide and water

→ Enthalpy change of formation conditions?

- One mole
- Standard states

→ Enthalpy change of atomisation conditions?

- One mole
- Standard conditions
- One mole of gaseous molecule forms (product)

→ What is Hess's law for combustion?



→ What is Hess's law for formation?



→ How do you use Hess's law to solve equations?

- Check whether its formation or combustion and draw accordingly
- Write the enthalpies for each product or reactant according to the arrows
- Keep in mind A is the elements or compounds used, B is the desired product, C is waste product

→ How to calculate enthalpy change?

- $Q = mct$, so $H = Q / n$

→ How to increase thermal stability?

- Decrease ionic charge
- Increase ionic radius

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Guillaume Favre



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Table of Contents October November 2chemistry Paper 60620

1. Understanding the eBook October November 2chemistry Paper 60620
 - The Rise of Digital Reading October November 2chemistry Paper 60620
 - Advantages of eBooks Over Traditional Books
2. Identifying October November 2chemistry Paper 60620
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an October November 2chemistry Paper 60620
 - User-Friendly Interface
4. Exploring eBook Recommendations from October November 2chemistry Paper 60620
 - Personalized Recommendations

- October November 2chemistry Paper 60620 User Reviews and Ratings
- October November 2chemistry Paper 60620 and Bestseller Lists
- 5. Accessing October November 2chemistry Paper 60620 Free and Paid eBooks
 - October November 2chemistry Paper 60620 Public Domain eBooks
 - October November 2chemistry Paper 60620 eBook Subscription Services
 - October November 2chemistry Paper 60620 Budget-Friendly Options
- 6. Navigating October November 2chemistry Paper 60620 eBook Formats
 - ePub, PDF, MOBI, and More
 - October November 2chemistry Paper 60620 Compatibility with Devices
 - October November 2chemistry Paper 60620 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of October November 2chemistry Paper 60620
 - Highlighting and Note-Taking October November 2chemistry Paper 60620
 - Interactive Elements October November 2chemistry Paper 60620
- 8. Staying Engaged with October November 2chemistry Paper 60620
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers October November 2chemistry Paper 60620
- 9. Balancing eBooks and Physical Books October November 2chemistry Paper 60620
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection October November 2chemistry Paper 60620
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine October November 2chemistry Paper 60620
 - Setting Reading Goals October November 2chemistry Paper 60620
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of October November 2chemistry Paper 60620
 - Fact-Checking eBook Content of October November 2chemistry Paper 60620

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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