

Formulae of Elements and Molecules

Complete the following table to show the structure and molecular formula of some covalent compounds. Remember: hydrogen and Group 7 elements form 1 covalent bond, oxygen forms 2, nitrogen forms 3, and carbon forms 4.

Name	Structure	Molecular formula
carbon dioxide		
		Cl ₂
ammonia		
butane		
		HCl

Formulae of Ionic Compounds

Complete the sentences using the words in the box. Use each word once.

negatively gain lose ions
positively electrostatic

Metal atoms _____ electrons and become _____ charged ions. Non metal atoms _____ electrons and become _____

charged ions. The compound formed has _____ bonding. The ions are held together by strong _____ forces.

Complete the table to show the formula and name for the compounds that could be made from each pair of ions. Think carefully, you must balance the charges. Use a periodic table.

Ion	Li ⁺	Na ⁺	Mg ²⁺	Ca ²⁺
F ⁻	LiF Lithium fluoride		MgF ₂ Magnesium fluoride	
O ²⁻				
Br ⁻				
Cl ⁻				
S ²⁻				

C3.1 - Chemical Reactions - Combined Chemistry Revision Sheet

Conservation of Mass

What is the Law of Conservation of mass?

What does it mean when we talk about a closed system?

If a system is not closed (open) why can the masses of substances appear to reduce during an experiment?

If a system is not closed (open) why can the masses of substances appear to increase during an experiment?

Calcium carbonate, CaCO₃, decomposes when heated. Calcium oxide, CaO, and carbon dioxide, CO₂, are made. Look at the equation for the reaction.



In an experiment, 20.0g of CaCO₃ are used. 11.2g of CaO are made. What is the mass of CO₂ made? g

Calculate the relative formula mass of calcium carbonate and calcium oxide (A_r values: C 12, O 16, Ca 40.1).

Calculate the mass of calcium oxide produced (to the nearest gram) when 50 g of calcium carbonate is heated.

Calculate the mass of carbon dioxide produced (to the nearest gram) when 200 g of calcium carbonate is heated.

Half Equations and Ionic Equations

Complete the sentences using the words in the box. Use each word once.

(aq) aqueous one don't covalent
ions water electrons both

Half equations model the change that happens to _____ reactant in a chemical reaction and shows how many _____ the atoms gain or lose. The balanced symbol equation for the reaction between sodium and chlorine is:



Write the two half equations for the sodium and chlorine ions and their electrons. Na _____ + _____ Cl₂ + _____

An ionic equation shows the _____ present in a reaction mixture. Soluble ionic compounds can be identified

in a solution as they have the _____

state symbol next to them. Ionic equations also include substances which are simple _____ bonded molecules. + g _____ Look at the following reaction

Identify the ions and write an ionic equation for the reaction.



_____ + _____ + _____ + _____ _____ + _____ + H₂O(l)

Spectator ions can be identified on both sides of the equation in the same form and do not change or take part in the reaction. Highlight them in the above equation. Write the net ionic equation for the ions that do change form.

Chemical Equations

Look at the following equations are they balanced or not balanced?

Count the number of each atom on either side. Write a 0 next to the ones which are balanced and fill next to the ones which are not balanced.



Chemical equation need to be balanced with the same number of each atoms going into the reaction as reactants as coming out of the reaction as products.

Balance the following equations.



When describing Fe₂O₃

chemical equations, we also add state symbols to show the state that a reactant or product is.

State symbol	Meaning
(s)	solid
(l)	liquid
(g)	gas
(aq)	aqueous

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Daniela Niemeyer



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