

Calculus-Based Thermodynamics (Practice Problems)

1. What will be the freezing-point and boiling-point of an aqueous solution containing 100 g of glucose, $C_6H_{12}O_6$, and 100 g of water? Equations: $K_f = 1.86^\circ C/m$, $K_b = 0.512^\circ C/m$.

$$\text{Molality} = \frac{\text{moles solute}}{\text{kg solvent}} = \frac{100 \text{ g} \times \frac{1 \text{ mol}}{180 \text{ g}}}{0.100 \text{ kg}} = 5.56 \text{ mol/kg}$$

$$\Delta T_f = K_f \cdot m = 1.86^\circ C/m \times 5.56 \text{ mol/kg} = 10.3^\circ C \quad \Delta T_b = K_b \cdot m = 0.512^\circ C/m \times 5.56 \text{ mol/kg}$$

2. What will be the freezing-point and boiling-point of an aqueous solution containing 100 g of NaCl and 100 g of water? Equations: $K_f = 1.86^\circ C/m$, $K_b = 0.512^\circ C/m$, $K_f(\text{NaCl}) = 3.72^\circ C/m$, $K_b(\text{NaCl}) = 1.02^\circ C/m$.

$$\text{Molality} = \frac{\text{moles solute}}{\text{kg solvent}} = \frac{100 \text{ g} \times \frac{1 \text{ mol}}{58.44 \text{ g}}}{0.100 \text{ kg}} = 1.71 \text{ mol/kg}$$

$$\Delta T_f = K_f \cdot m = 3.72^\circ C/m \times 1.71 \text{ mol/kg} = 6.36^\circ C \quad \Delta T_b = K_b \cdot m = 1.02^\circ C/m \times 1.71 \text{ mol/kg}$$

3. A 100 mL sample of 0.10 M $NaCl$ is cooled to $-10^\circ C$ and the freezing-point depression is determined to be $0.34^\circ C$. Assume that the solute does not freeze. Equations: $K_f = 1.86^\circ C/m$.

$$\Delta T_f = K_f \cdot m \quad 0.34^\circ C = 1.86^\circ C/m \cdot m \quad m = 0.183 \text{ mol/kg}$$

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$$\Rightarrow m = 0.183 \text{ mol/kg}$$

$$m = 0.183 \text{ mol/kg}$$

$$m = \frac{\text{moles solute}}{\text{kg solvent}} \quad 0.183 \text{ mol/kg} = \frac{x}{0.100 \text{ kg}}$$

$$\Rightarrow x = 0.0183 \text{ mol} = 0.0183 \text{ mol} \times 58.44 \text{ g/mol}$$

$$\text{Mass} = 1.06 \text{ g}$$

$$\text{Total mass} = 100 \text{ g} + 1.06 \text{ g} = 101.06 \text{ g} \quad \text{Mass} = 1.06 \text{ g}$$

4. A solution composed from 10 g of glucose, $C_6H_{12}O_6$, and 100 g of water is cooled. The freezing-point depression is $0.52^\circ C$. The normal freezing point of pure water is $0^\circ C$ and the freezing-point depression constant for water is $1.86^\circ C/m$. Determine the molar mass of the compound. Equations: $K_f = 1.86^\circ C/m$, $K_b = 0.512^\circ C/m$.

$$\Delta T_f = K_f \cdot m \quad 0.52^\circ C = 1.86^\circ C/m \cdot m \quad m = 0.28 \text{ mol/kg}$$

$$m = \frac{\text{moles solute}}{\text{kg solvent}} \quad 0.28 \text{ mol/kg} = \frac{x}{0.100 \text{ kg}}$$

$$\Rightarrow x = 0.028 \text{ mol} = 0.028 \text{ mol} \times M$$

$$M = 182 \text{ g/mol}$$

$$m = \frac{\text{moles solute}}{\text{kg solvent}} \quad 0.28 \text{ mol/kg} = \frac{x}{0.100 \text{ kg}}$$

$$\Rightarrow x = 0.028 \text{ mol} = 0.028 \text{ mol} \times M$$

$$\text{Mass} = 182 \text{ g/mol}$$

$$\Delta T_f = K_f \cdot m \quad 0.52^\circ C = 1.86^\circ C/m \cdot m \quad m = 0.28 \text{ mol/kg}$$

$$\text{Mass} = 182 \text{ g/mol}$$

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