

Stoichiometry Review

1. The combustion of propane occurs via the reaction:



How many grams of oxygen are required to burn completely 10.0 g of propane?

2. Small quantities of chlorine can be prepared in the laboratory by the reaction:



What volume of chlorine at STP can be prepared from 100.0 g of manganese dioxide?

3. The most common ore of arsenic is mispickel, FeSAs . Upon heating this ore, free arsenic is obtained. How many grams of FeSAs are required to produce 10.0 g of arsenic according to the following reaction?



4. Glucose is used as an energy source by the human body. The overall reaction in the body is:



Calculate the volume of oxygen at STP required to convert 28.0 g of glucose to carbon dioxide and water.

5. Potassium nitrate is widely used as a fertilizer because it provides two essential elements, potassium and nitrogen. It is made by mixing potassium chloride and nitric acid in the presence of oxygen according to the equation:



How many kilograms of potassium nitrate will be produced from 50.0 kg of potassium chloride and 50.0 kg of nitric acid with excess oxygen?

6. Bromine can be prepared by adding chlorine to an aqueous solution of sodium bromide. How many grams of bromine are formed if 25.0 g of chlorine and 25.0 g of sodium bromide are reacted according to the reaction below? Which reactant is in excess and how much remains after the reaction is complete?



7. How many millilitres of 2.00 M hydrochloric acid are required to react with 2.55 g of zinc according to the following reaction:



8. A 0.473 g sample of phosphorus is reacted with an excess of chlorine. Calculate the mass of phosphorus pentachloride produced.

9. How many grams of oxygen can be prepared by the decomposition of 25.0 g of mercury(II) oxide?

10. What volume of oxygen gas at STP can be prepared by the decomposition of 25.0 g of potassium chlorate?

11. How many grams of zinc are required for the replacement of 0.100 g of hydrogen from sulfuric acid?

12. What mass of magnesium chloride is required to react with 10.0 g of silver nitrate?

Review Stoichiometry Section 2 Answers With Work

**Thirumalaisamy P.
Velavancorresponding**



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