

b) How many grams of phosphorus are required to react completely with 20.0 g of  $O_2$ ?

$$20.0 \text{ g } O_2 \times \frac{1 \text{ mol } O_2}{31.998 \text{ g } O_2} \times \frac{4 \text{ mol P}}{5 \text{ mol } O_2} \times \frac{30.973 \text{ g P}}{1 \text{ mol P}} = 15.5 \text{ g P}$$

c) What is the theoretical yield in grams if you combine the amounts of reactants in part b?

$$20.0 \text{ g } O_2 \times \frac{1 \text{ mol } O_2}{31.998 \text{ g } O_2} \times \frac{2 \text{ mol } P_2O_5}{5 \text{ mol } O_2} \times \frac{141.941 \text{ g } P_2O_5}{1 \text{ mol } P_2O_5} = 35.5 \text{ g } P_2O_5$$

OR

$$15.5 \text{ g P} \times \frac{1 \text{ mol P}}{30.973 \text{ g P}} \times \frac{2 \text{ mol } P_2O_5}{4 \text{ mol P}} \times \frac{141.941 \text{ g } P_2O_5}{1 \text{ mol } P_2O_5} = 35.5 \text{ g } P_2O_5$$

5. A gaseous mixture containing 5.00 moles of  $H_2$  and 7.00 moles of  $Br_2$  reacts to form  $HBr$ .

a) Write a balanced chemical equation for this reaction.



b) Which reactant is limiting?

1 to 1 ratio  
5.00 mol  $H_2$  & 7.00 mol  $Br_2$   
 $\therefore H_2$  is limiting

c) What is the theoretical yield for this reaction in moles?

$$5.00 \text{ mol } H_2 \times \frac{2 \text{ mol } HBr}{1 \text{ mol } H_2} = 10.0 \text{ mol } HBr$$

d) What is the theoretical yield for this reaction in grams?

$$10.0 \text{ mol } HBr \times \frac{80.912 \text{ g } HBr}{1 \text{ mol } HBr} = 809 \text{ g } HBr$$

6. Sodium reacts with molecular hydrogen to form sodium hydride. A reaction mixture contains 10.0 g Na and 0.0235 g  $H_2$ .

a) Write a balanced chemical equation for this reaction.



b) Which reactant is limiting?  $H_2$  is limiting

$$10.00 \text{ g Na} \times \frac{1 \text{ mol Na}}{22.9898 \text{ g Na}} = 0.4350 \text{ mol Na} \quad 0.0235 \text{ g } H_2 \times \frac{1 \text{ mol } H_2}{2.0158 \text{ g } H_2} = 0.0117 \text{ mol } H_2$$
$$\frac{0.4350}{2} = 0.2175 \quad \frac{0.0117}{1} = 0.0117$$

c) What is the theoretical yield for this reaction in grams?

$$0.0117 \text{ mol } H_2 \times \frac{2 \text{ mol NaH}}{1 \text{ mol } H_2} \times \frac{23.998 \text{ g NaH}}{1 \text{ mol NaH}} = 0.562 \text{ g NaH}$$

d) When this reaction is actually performed, 0.428 g of NaH is recovered. What is the percent yield of the reaction?

$$\frac{0.428 \text{ g}}{0.562} \times 100 = 76.2\%$$

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**SA Dillow**



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## **Table of Contents Periodic Trends Pogil Extension Questions Key**

1. Understanding the eBook Periodic Trends Pogil Extension Questions Key
  - The Rise of Digital Reading Periodic Trends Pogil Extension Questions Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Trends Pogil Extension Questions Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Periodic Trends Pogil Extension Questions Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Trends Pogil Extension Questions Key
  - Personalized Recommendations
  - Periodic Trends Pogil Extension Questions Key User Reviews and Ratings
  - Periodic Trends Pogil Extension Questions Key and Bestseller Lists
5. Accessing Periodic Trends Pogil Extension Questions Key Free and Paid eBooks
  - Periodic Trends Pogil Extension Questions Key Public Domain eBooks
  - Periodic Trends Pogil Extension Questions Key eBook Subscription Services
  - Periodic Trends Pogil Extension Questions Key Budget-Friendly Options
6. Navigating Periodic Trends Pogil Extension Questions Key eBook Formats

- ePub, PDF, MOBI, and More
- Periodic Trends Pogil Extension Questions Key Compatibility with Devices
- Periodic Trends Pogil Extension Questions Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Periodic Trends Pogil Extension Questions Key
  - Highlighting and Note-Taking Periodic Trends Pogil Extension Questions Key
  - Interactive Elements Periodic Trends Pogil Extension Questions Key
- 8. Staying Engaged with Periodic Trends Pogil Extension Questions Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Periodic Trends Pogil Extension Questions Key
- 9. Balancing eBooks and Physical Books Periodic Trends Pogil Extension Questions Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Periodic Trends Pogil Extension Questions Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Periodic Trends Pogil Extension Questions Key
  - Setting Reading Goals Periodic Trends Pogil Extension Questions Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Periodic Trends Pogil Extension Questions Key
  - Fact-Checking eBook Content of Periodic Trends Pogil Extension Questions Key
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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