

The Periodic Law

SECTION 2

Select the correct answer for the following questions for the space provided.

1. _____ for the modern periodic table, elements are ordered
 - (a) according to increasing atomic mass
 - (b) according to Mendeleev's original design
 - (c) according to increasing atomic number
 - (d) based on how they were discovered
2. _____ Mendeleev's original short periodic table gave useful chemical properties of elements appeared in regular intervals and also predicted elements were missing because of his saying
 - (a) theory
 - (b) reasoning
 - (c) atomic number
 - (d) atomic mass
3. _____ The modern periodic law states that
 - (a) all new elements with the same properties are listed in the same group every year
 - (b) the physical and chemical properties of elements are functions of atomic number
 - (c) elements exhibit properties of both periods and groups
 - (d) the chemical properties of elements can be grouped according to periodicity but physical properties cannot
4. _____ The discovery of the noble gases changed Mendeleev's periodic table
 - (a) adding a row
 - (b) adding a new period
 - (c) adding a new column
 - (d) adding a new group
5. _____ The main reason the properties of the noble gases is that they are
 - (a) inactive
 - (b) unreactive
 - (c) large atoms
 - (d) large gaseous molecules
6. _____ Likewise, the first column in Group 1, has an atomic number of 3. The second element in this group has an atomic number of
 - (a) 4
 - (b) 5
 - (c) 6
 - (d) 7
7. How groups in Mendeleev's periodic table of 18 and the atomic number of 55
 - (a) 1. How many periods are in this series?
 - (b) 2. How many elements are in this group?
 - (c) 3. What is the element symbol for this element? How is it named? How does it compare to the other elements in this group?

Periodic Law Chapter 5 Test Answers

**Norman E. Griswold, Henry Fuller
Holtzclaw**



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