

Chapter 1: Limiting Reagents and Percent Yield (Pages 1-4)

Remember: Limiting reagent is the reactant that will completely react and leave another reactant

Excess: All reactants except the limiting reagent will be left over. Remember, limiting reagent is the reactant that will completely react.

Remember, limiting reagent will determine how much product is produced and byproducts will be the leftover reactant in a chemical reaction.



This reaction will work completely for 2 molecules of H_2 .

$2\text{H}_2(\text{g})$	200.00 grams
$\text{O}_2(\text{g})$	64.00 grams
$2\text{H}_2\text{O}(\text{l})$	180.00 grams
$\text{H}_2\text{O}(\text{l})$	90.00 grams

- (1) 200.00 g of H_2 will completely react with 64.00 g of O_2 to form 180.00 g of H_2O .
Remember, limiting reagent will determine how much product is produced.



- (2) $\frac{200.00 \text{ g}}{2.016 \text{ g}} = 99.206$ (the limiting reagent is O_2) for this reaction.

- (3) 180.00 g of H_2O will be produced from 200.00 g of H_2 and 64.00 g of O_2 .



- (4) 180.00 g of H_2O will be produced from 200.00 g of H_2 and 64.00 g of O_2 .
Remember, limiting reagent will determine how much product is produced.



Percent Yield And Limiting Reactants Quiz

Purshotam Kaushik & Kirti Kaushik

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the author names.

Percent Yield And Limiting Reactants Quiz:

Whispering the Secrets of Language: An Emotional Journey through **Percent Yield And Limiting Reactants Quiz**

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