

Periodic Trends Worksheet

Directions: Use your notes to answer the following questions.

1. Rank the following elements by increasing atomic radius: carbon, aluminum, oxygen, potassium.

Oxygen < Carbon < Aluminum < Potassium

2. Rank the following elements by increasing electronegativity: sulfur, oxygen, neon, aluminum.

Neon < Aluminum < Sulfur < Oxygen

3. Why does fluorine have a higher ionization energy than iodine?

Fluorine has nine protons attracting 9 electrons which are much closer to the nucleus compared to iodine. This results in smaller size of Fluorine than iodine making it difficult to remove an electron from the atom.

4. Why do elements in the same family generally have similar properties?

Because they have same number of electrons in the outer shell (valence electrons) which take part in chemical reaction.

5. Indicate whether the following properties increase or decrease from left to right across the periodic table.

- | | |
|--|------------------|
| a. atomic radius (excluding noble gases) | Decreases |
| b. first ionization energy | Increases |
| c. electronegativity | Increases |

6. What trend in atomic radius occurs down a group on the periodic table? What causes this trend?

Atomic radius **decreases** down the group on the periodic table. As we compare the elements down the group, the effective nuclear charge increases, but at the same time the outermost electrons are found in the shell that is farther away from the nucleus. Also the number of screening electrons increases. This results in reduced attraction between the nucleus and the outermost electrons.

7. What trend in ionization energy occurs across a period on the periodic table? What causes this trend?

Ionization energy **increases** from left to right across the period. The atomic size becomes smaller from left to right. So it becomes harder to remove electron from the atom. Hence the energy required to do so (ionization Energy) increases.

Periodic Table Trend Activities Answer Key

The Princeton Review



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Chemistry in the Community. American Chemical Society, 2002 This volume has relevance to a wide number of courses giving a hands on introduction to chemistry in relation to community issues rather than around specific chemical concepts.

Chemistry in the Community (Enhanced Core Four) American Chemical Society, 2006-02-15

Problems and Problem Solving in Chemistry Education Georgios Tsapalis, 2021-05-17 Problem solving is central to the teaching and learning of chemistry at secondary, tertiary and post tertiary levels of education opening to students and professional chemists alike a whole new world for analysing data, looking for patterns and making deductions. As an important higher order thinking skill, problem solving also constitutes a major research field in science education. Relevant education research is an ongoing process with recent developments occurring not only in the area of quantitative computational problems but also in qualitative problem solving. The following situations are considered: some general others with a focus on specific areas of chemistry: quantitative problems, qualitative reasoning, metacognition and resource activation, deconstructing the problem solving process, an overview of the working memory hypothesis, reasoning with the electron pushing formalism, scaffolding organic synthesis skills, spectroscopy for structural characterization in organic chemistry, enzyme kinetics, problem solving in the academic chemistry laboratory, chemistry problem solving in context, team based active learning, technology for molecular representations, IR spectra simulation and computational quantum chemistry tools. The book concludes with methodological and epistemological issues in problem solving research and other perspectives in problem solving in chemistry. With a foreword by George Bodner.

Science Units for Grades 9-12 Randy L. Bell, Joe Garofalo, 2005

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Trends in Teaching Experimentation in the Life Sciences Nancy J. Pelaez, Stephanie M. Gardner, Trevor R. Anderson, 2022-05-11 This book is a guide for educators on how to develop and evaluate evidence based strategies for teaching biological experimentation to thereby improve existing and develop new curricula It unveils the flawed assumptions made at the classroom department and institutional level about what students are learning and what help they might need to develop competence in biological experimentation Specific case studies illustrate a comprehensive list of key scientific competencies that unpack what it means to be a competent experimental life scientist It includes explicit evidence based guidelines for educators regarding the teaching learning and assessment of biological research competencies The book also provides practical teacher guides and exemplars of assignments and assessments It contains a complete analysis of the variety of tools developed thus far to assess learning in this domain This book contributes to the growth of public understanding of biological issues including scientific literacy and the crucial importance of evidence based decision making around public policy It will be beneficial to life science instructors biology education researchers and science administrators

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