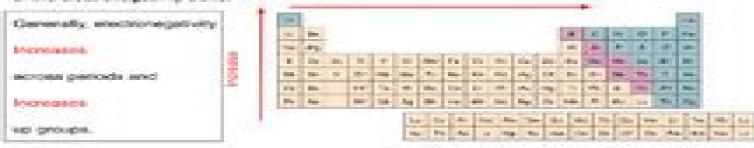
Periodicity - Electronegativity

Name

Angeler Key

 Fill in the blanks below, and then ones arrows on to the periodic table indicating the direction of the electronogativity trend.



List the elements between cross of increasing electronegativity. (Smallest to largest).

Ba, Ra, Ca, Be, Mg	Zn. Nt. Gu, Sc. Ti	Ft. Bt. Ct. Pb. Fr
Ro, Ba, Ca, Mg, Be	de, B. Nt. Gu, Zn	Ft. RD. B. Ct. F
No. Sn. Rb. Fr. I	F, LJ, Be, K, Ma	P. St. Ar, Mg. Co.
Fr. Rb. Nb. Sn. I	K, No. LJ, Ser. F	Co. Mg. St. P. Ar

Are the elements below in order of decreasing electronegativity? If year, write "yes". If not, write them in the correct order.

Cir. K. Ser. Br Sir, Ser. Cir. K.	E. Bo. CL F F. Cl. Bo. (O. B. Li, No.
Y. Se, Rb. Fr	All, Sc. Sr. Bo	Au. Pb. Te. S S. Te. Ph. Au

Briefly explain why electronogativity changes up groups.

Generally, atactronogistivity increases as you move up in a group. This is because the elemic number decreases, and as a result, there is less discarce between the nucleus and valence excitents. In other words, there is a smaller atomic radius.

5. Shiefly explain why electronegativity changes across periods.

As you move than left to right, across a period, the nuclear charge of elements increases faster than electron shielding. This results in an increased attraction to electrons.

Periodicity Worksheet Answer Guide

TD Snyder

Periodicity Worksheet Answer Guide:

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