

## CLO- Classification of Elements & Periodicity in Properties

### MULTIPLE CHOICE QUESTIONS:

Choose the one correct option in the following:

- Identify the least stable ion amongst the following :  
(a)  $\text{Li}^+$  (b)  $\text{Be}^+$  (c)  $\text{B}^+$  (d)  $\text{C}^+$
- The correct order of radii is :  
(a)  $\text{N} < \text{Be} < \text{B}$  (b)  $\text{F}^- < \text{O}^{2-} < \text{N}^{3-}$  (c)  $\text{N} < \text{Li} < \text{K}$  (d)  $\text{Fe}^{2+} < \text{Fe}^{3+} < \text{Fe}^{4+}$
- Which one of the following has largest size :  
(a)  $\text{Na}$  (b)  $\text{Na}^+$  (c)  $\text{Na}^-$  (d)  $\text{Li}^+$
- The number of elements in the 4<sup>th</sup> period of periodic table is :  
(a) 8 (b) 10 (c) 18 (d) 32
- Which of the following has least electron affinity (electron gain enthalpy) :  
(a)  $\text{O}$  (b)  $\text{N}$  (c)  $\text{Ar}$  (d)  $\text{F}$
- An element of atomic number 29 belongs to which of the following block of the periodic table :  
(a) s-block (b) p-block (c) d-block (d) f-block
- Which one of the following arrangements not truly represents the property indicated against it:  
(a)  $\text{Br}_2 < \text{Cl}_2 < \text{F}_2$  – electronegativity (b)  $\text{Br}_2 < \text{F}_2 < \text{Cl}_2$  – Electron Affinity  
(c)  $\text{Br}_2 < \text{Cl}_2 < \text{F}_2$  – Bond energy (d)  $\text{Br}_2 < \text{Cl}_2 < \text{F}_2$  – Oxidising Power
- Alkali metals in each period have :  
(a) Smallest size (b) Lowest IE (c) Highest IE (d) Highest EN
- A sudden large jump between the value of second and third ionization energies of elements would be associated with which of the following electronic configuration :  
(a)  $2\text{p}^4 3\text{s}^1$  (b)  $3\text{s}^2 3\text{p}^1$  (c)  $3\text{s}^2 3\text{p}^2$  (d)  $2\text{p}^4 3\text{s}^2$
- Which of the following pairs are chemically dissimilar :  
(a)  $\text{Na}$  and  $\text{K}$  (b)  $\text{Ba}$  and  $\text{Sr}$  (c)  $\text{Zr}$  and  $\text{Hf}$  (d)  $\text{Ca}$  and  $\text{Zn}$
- Collective name given to the elements with outer shell configuration  $ns^2 np^6$  is :  
(a) Chalkogens (b) Alkaline earths metals (c) Transition metals (d) Nobel gases
- Modern extended periodic table was given by :  
(a) Bohr (b) Lothar mayer (c) Moseley (d) None alone
- The most electronegative elements is :  
(a) Hydrogen (b) Fluorine (c) Chlorine (d) Oxygen
- The element having electronic configuration :  
 $[\text{Kr}]4\text{d}^{10}4\text{f}^{14}5\text{s}^2 5\text{p}^6 5\text{d}^1 6\text{s}^2$  belongs to  
(a) s-block (b) p-block (c) d-block (d) f-block
- An elements R forms the highest oxide  $\text{R}_2\text{O}_5$ . R belong to :

# Periodicity Multiple Choice Questions

**Thomas Griffiths**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

## **Periodicity Multiple Choice Questions:**

## Decoding **Periodicity Multiple Choice Questions**: Revealing the Captivating Potential of Verbal Expression

In an era characterized by interconnectedness and an insatiable thirst for knowledge, the captivating potential of verbal expression has emerged as a formidable force. Its ability to evoke sentiments, stimulate introspection, and incite profound transformations is genuinely awe-inspiring. Within the pages of "**Periodicity Multiple Choice Questions**," a mesmerizing literary creation penned with a celebrated wordsmith, readers set about an enlightening odyssey, unraveling the intricate significance of language and its enduring affect our lives. In this appraisal, we shall explore the book's central themes, evaluate its distinctive writing style, and gauge its pervasive influence on the hearts and minds of its readership.

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