

10. Circle the element in each of the following pairs with the greater (1st) ionization energy:

a) lithium or boron

c) cesium or lead

b) magnesium or strontium

d) sodium or argon

11. Arrange the following elements in order of increasing ionization energy:

a) Sr, Be, Mg

b) Bi, Cs, Ba

c) Na, Al, S

12. Why is there a very large increase between the 1st and 2nd ionization energies for the elements in group 1?

Name: _____ Date: _____ Class period: _____



Unit: Periodic Trends "Ionic Size and E.N. Trend" - WS #4

Directions: Please answer each fill in the blank with the best answer. ☺

1. A cation is an ion with an overall _____ charge.
2. Cations are made when an atom _____.
3. The size of a positive ion is always _____ than the size of the neutral atom of the same element.
4. As the positive charge on an ion increases, the size of the atom will _____.
5. An anion is an ion with an overall _____ charge.
6. Anions are made when at atom _____.
7. The size of a negative ion is always _____ than the size of the neutral atom of the same element.
8. As the negative charge on an ion increases, the size of the atom will _____.
9. Which element/ion has the **biggest** atomic/ionic size in each of the following pairs? Please circle your answer.
 - a) Na or Na⁺
 - b) S or S²⁻
 - c) O⁻¹ or O⁻²
 - d) Al or Al⁺³
 - e) Cu⁺¹ or Cu⁺²
10. _____ is the tendency for the atoms of the element to attract electrons when they are chemically combined with another element.

Periodic Trends Unit Exam And Answers

YCT Expert Team



Periodic Trends Unit Exam And Answers:

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