

Annotated periodic table of the elements

1		2		Group name*																		13		14		15		16		17		18																																																					
1 H 1.01 Hydrogen				1 Hydrogen & the alkali metals 2 Alkaline earth metals 11 Coinage metals 12 Volatile metals 13 Icosagens																		14 Crystallogens 15 Pnictogens 16 Chalcogens 17 Halogens 18 Helium group		3 Li 6.94 Lithium		4 Be 9.01 Beryllium		5 B 10.81 Boron		6 C 12.01 Carbon		7 N 14.01 Nitrogen		8 O 16.00 Oxygen		9 F 18.99 Fluorine		10 Ne 20.18 Neon																																															
3 Li 6.94 Lithium		4 Be 9.01 Beryllium		*Groups 3 to 10 are named after their first members i.e. Group 3 is the Scandium group																		11 Na 22.99 Sodium		12 Mg 24.31 Magnesium		13 Al 26.98 Aluminium		14 Si 28.09 Silicon		15 P 30.97 Phosphorus		16 S 32.06 Sulfur		17 Cl 35.45 Chlorine		18 Ar 39.95 Argon																																																	
11 Na 22.99 Sodium		12 Mg 24.31 Magnesium																				19 K 39.10 Potassium		20 Ca 40.08 Calcium		21 Sc 44.96 Scandium		22 Ti 47.87 Titanium		23 V 50.94 Vanadium		24 Cr 52.00 Chromium		25 Mn 54.94 Manganese		26 Fe 55.85 Iron		27 Co 58.93 Cobalt		28 Ni 58.69 Nickel		29 Cu 63.55 Copper		30 Zn 65.38 Zinc		31 Ga 69.72 Gallium		32 Ge 72.64 Germanium		33 As 74.92 Arsenic		34 Se 78.96 Selenium		35 Br 79.90 Bromine		36 Kr 83.80 Krypton																													
37 Rb 85.47 Rubidium		38 Sr 87.62 Strontium																				39 Y 88.91 Yttrium		40 Zr 91.22 Zirconium		41 Nb 92.91 Niobium		42 Mo 95.94 Molybdenum		43 Tc 98.91 Technetium		44 Ru 101.07 Ruthenium		45 Rh 101.07 Rhodium		46 Pd 106.32 Palladium		47 Ag 107.87 Silver		48 Cd 112.41 Cadmium		49 In 114.82 Indium		50 Sn 118.71 Tin		51 Sb 121.76 Antimony		52 Te 127.60 Tellurium		53 I 126.91 Iodine		54 Xe 131.29 Xenon																																	
55 Cs 132.91 Cesium		56 Ba 137.33 Barium		57 to 70																		71 Lu 174.97 Lutetium		72 Hf 178.49 Hafnium		73 Ta 180.95 Tantalum		74 W 183.84 Tungsten		75 Re 186.21 Rhenium		76 Os 190.23 Osmium		77 Ir 192.22 Iridium		78 Pt 195.08 Platinum		79 Au 196.97 Gold		80 Hg 200.59 Mercury		81 Tl 204.38 Thallium		82 Pb 207.2 Lead		83 Bi 208.98 Bismuth		84 Po 209 Polonium		85 At 210 Astatine		86 Rn 222 Radon																																	
87 Fr 223 Francium		88 Ra 226 Radium		89 to 102																		103 La 138.91 Lanthanum		104 Ce 140.12 Cerium		105 Pr 140.91 Praseodymium		106 Nd 144.24 Neodymium		107 Pm 144.91 Promethium		108 Sm 150.36 Samarium		109 Eu 151.96 Europium		110 Gd 157.25 Gadolinium		111 Tb 158.93 Terbium		112 Dy 162.50 Dysprosium		113 Ho 164.93 Holmium		114 Er 167.26 Erbium		115 Tm 168.93 Thulium		116 Yb 173.05 Ytterbium		117 Lu 174.97 Lutetium		118 Hf 178.49 Hafnium		119 Ta 180.95 Tantalum		120 W 183.84 Tungsten		121 Re 186.21 Rhenium		122 Os 190.23 Osmium		123 Ir 192.22 Iridium		124 Pt 195.08 Platinum		125 Au 196.97 Gold		126 Hg 200.59 Mercury		127 Tl 204.38 Thallium		128 Pb 207.2 Lead		129 Bi 208.98 Bismuth		130 Po 209 Polonium		131 At 210 Astatine		132 Rn 222 Radon					
113 In 114.82 Indium		114 Sn 118.71 Tin																				119 Cs 132.91 Cesium		120 Ba 137.33 Barium		121 La 138.91 Lanthanum		122 Ce 140.12 Cerium		123 Pr 140.91 Praseodymium		124 Nd 144.24 Neodymium		125 Pm 144.91 Promethium		126 Sm 150.36 Samarium		127 Eu 151.96 Europium		128 Gd 157.25 Gadolinium		129 Tb 158.93 Terbium		130 Dy 162.50 Dysprosium		131 Ho 164.93 Holmium		132 Er 167.26 Erbium		133 Tm 168.93 Thulium		134 Yb 173.05 Ytterbium		135 Lu 174.97 Lutetium		136 Hf 178.49 Hafnium		137 Ta 180.95 Tantalum		138 W 183.84 Tungsten		139 Re 186.21 Rhenium		140 Os 190.23 Osmium		141 Ir 192.22 Iridium		142 Pt 195.08 Platinum		143 Au 196.97 Gold		144 Hg 200.59 Mercury		145 Tl 204.38 Thallium		146 Pb 207.2 Lead		147 Bi 208.98 Bismuth		148 Po 209 Polonium		149 At 210 Astatine		150 Rn 222 Radon	

- Group names*
- 1 Hydrogen & the alkali metals
 - 2 Alkaline earth metals
 - 11 Coinage metals
 - 12 Volatile metals
 - 13 Lanthanoids
 - 14 Chalcogens
 - 15 Pnictogens
 - 16 Chalcogens
 - 17 Halogens
 - 18 Helium group

*Groups 3 to 10 are named after their first members i.e. Group 3 is the Scandium group

CLASS COUNTERPARTS

Metals	Nonmetals
Pre-transition	
Active	Active
Transition	Intermediate
Post-transition	Metalloids
Metals (Ru-Pt, Os-Au)	Metals (Ru-Pt, Os-Au)

57 La 138.91 Lanthanum	58 Ce 140.12 Cerium	59 Pr 140.91 Praseodymium	60 Nd 144.24 Neodymium	61 Pm 144.91 Promethium	62 Sm 150.36 Samarium	63 Eu 151.96 Europium	64 Gd 157.25 Gadolinium	65 Tb 158.93 Terbium	66 Dy 162.50 Dysprosium	67 Ho 164.93 Holmium	68 Er 167.26 Erbium	69 Tm 168.93 Thulium	70 Yb 173.05 Ytterbium
89 Ac 227.03 Actinium	90 Th 232.04 Thorium	91 Pa 231.04 Protactinium	92 U 238.03 Uranium	93 Np 237.05 Neptunium	94 Pu 244.06 Plutonium	95 Am 243.06 Americium	96 Cm 247.07 Curium	97 Bk 247.07 Berkelium	98 Cf 251.08 Californium	99 Es 252.08 Einsteinium	100 Fm 257.10 Fermium	101 Md 258.10 Mendelevium	102 No 259.10 Nobelium

Radioactive

Oganesson is likely to be either a metalloid or a post-transition metal

Active metals

Pre-transition metals
Light, reactive, mostly of low mechanical strength, melting and boiling points. The group 1 members are easily cut with a knife. Only Be, Mg and Al have any structural uses. Most form basic oxides (Be and Al are amphoteric).

Rare earths
Closely resemble Ca, Sr and Ba, but many are heavier and they form mostly pale coloured compounds. Similar to one another, esp. La to Lu (aka the lanthanides), and hard to separate for that reason. Abundant but widely dispersed; commercially viable concentrations are rare. Oxides are strongly to moderately basic. Sc, Yb and Lu exhibit weak amphotericism.

Actinides
Soft, dense and reactive metals. Those from Ra onwards are synthetic. Early actinides (Th-Am) show some similarities to transition metals and have basic, amphoteric, or acidic oxides; late actinides are more like lanthanides. Many actinides form coloured compounds.

Transition metals
Moderate to high density metals, with high melting and boiling points, many have high hardness and mechanical strength. Those from Rf onwards are synthetic. Chemically show variable valency and strong tendency to form coordination and colored compounds. Oxides are basic, amphoteric or acidic, depending on oxidation state.

Noble metals
Resistant to corrosion and oxidation, unlike base metals such as Fe, Ni, Cu and Zn. Commonly found together in native form. Good catalytic properties. Very weakly electropositive; insoluble oxides.

Frontier metals
Soft (or brittle), low strength, and with melting points lower than transition metals. Crystalline structures tend to show directional bonding, with generally greater complexity or lower nearest neighbours than other metals. All used or tried as soft metal lubricants. Most are chemically weak; a minority are moderately electropositive. Characterised, to varying degrees, by covalent bonding tendencies, amphotericism, and anionic species such as zincates, stannates, and bismuthates.

Metalloids
Metallic looking brittle, fair electrical conductors. All known in semiconducting forms. Moderate ionisation energies (750–1,000 kJ/mol), and electronegativity values (1.8–2.3). Mostly behave chemically as non-metals. Can form alloys with metals. Amphoteric or weakly acidic oxides.

Intermediate nonmetals
H and N are colourless gases; C, P, and Se are metalloids looking solids while S is yellow. The latter four are brittle but all known in malleable, plastic or ductile forms. Moderate to high ionisation energies, low to high electron affinities, moderate to high electronegativity values, relatively poor to moderately strong oxidising agents, tend to form covalent or metallic to near-metallic compounds with metals. Their oxides are acidic, or neutral (H₂O, CO, NO, N₂O).

*as black P, the most stable form in ambient conditions.

Group 1
The alkali metals Li, Na, K, Rb, Cs, and Fr are soft, silvery-white metals. They are highly reactive, especially with water, and form ionic compounds. They are found in Group 1 of the periodic table.

Covalent nonmetals
O, F, and Cl are, respectively, colourless, pale yellow, and pale yellow-green gases. Br is a red-brown liquid. I is a metallic looking solid. Chemically tend to have high ionisation energies, electron affinities, and electronegativity values, and be relatively strong oxidising agents. Manifestations include oxygen's major association with the ubiquitous processes of corrosion and combustion, and the intrinsically covalent nature of the non-metallic halogens. All five tend to form predominantly ionic compounds with metals. Their oxides are acidic.

Noble gases
Colourless, odourless, non-flammable gases with very low chemical reactivity. The first noble gas compound, of approximate composition XePtF₆, was not prepared until 1962. More than half-a-thousand noble gas compounds were synthesised in the ensuing half-century. Compounds of He, Ne, and Ar, in ambient conditions, are not known.

Periodic Table Guide Answer Sheet

Robert R. Redfield



Periodic Table Guide Answer Sheet:

Whispering the Techniques of Language: An Emotional Quest through **Periodic Table Guide Answer Sheet**

In a digitally-driven world where displays reign great and instant conversation drowns out the subtleties of language, the profound secrets and psychological subtleties hidden within words often get unheard. However, set within the pages of **Periodic Table Guide Answer Sheet** a charming literary value pulsing with fresh thoughts, lies an exceptional quest waiting to be undertaken. Written by a talented wordsmith, this marvelous opus encourages visitors on an introspective trip, lightly unraveling the veiled truths and profound influence resonating within the muscles fabric of each word. Within the psychological depths of this emotional review, we will embark upon a sincere exploration of the book's primary subjects, dissect its fascinating writing model, and yield to the effective resonance it evokes strong within the recesses of readers' hearts.

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