# Graphing Periodic Trends Lab

Name\_\_\_\_\_Hr\_\_\_

PURPOSE: To determine how certain properties are periodic when the elements are arranged in periods or groups.

DISCUSSION: There are many periodic trends that occur as you cross a period or go down a group/family on the Periodic Table. Graphing various elemental properties will help to illustrate these trends.

#### REFERENCE:

Angstrom - A unit of distance equal to one ten-billionth of a meter.

Atomic Radius - The distance from the center of the nucleus to the valence electron energy level.

ionization energy - The energy required to remove an electron from an atom.

#### DATA

Properties of the First 18 Elements

Properties of the First 18		Elements	
Element	Atomic Number	Atomic Radius (Angstroms)	Ionization Energy
н	1	0.79	314
He	2	0.49	567
Li	3	2.05	124
Ве	4	1.4	215
8	5	1.17	191
C	6	0.91	260
M	7	0.75	335
0	8	0.65	314
F	9	0.57	402
No	10	7	2
Ma	1.1	2.23	119
Mg	12	1.72	176
A.I	13	1.82	138
Si	1.4	1.46	188
P	15	1.23	242
S	16	1.09	239
CI	17	0.97	299
Ar	18	?	2

Using the data above, make a line graph the following data. Draw a dark vertical line to separate the 3 periods. Give each graph a title and label both the x-axis and the y-axis. Connect your points. Predict the plot points for elements with guestions marks.

- Atomic Radius (y) vs. Atomic Number (x) for the first 3 periods (18 elements 3 graphs)
- 2. Ionization energy (y) vs. Atomic Number (x) for the first 3 periods (18 elements 3 graphs)

## **Periodic Trends Lab Answers**

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#### **Table of Contents Periodic Trends Lab Answers**

- 1. Understanding the eBook Periodic Trends Lab Answers
  - The Rise of Digital Reading Periodic Trends Lab Answers
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Periodic Trends Lab Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Periodic Trends Lab Answers
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Periodic Trends Lab Answers
  - Personalized Recommendations
  - Periodic Trends Lab Answers User Reviews and Ratings

- Periodic Trends Lab Answers and Bestseller Lists
- 5. Accessing Periodic Trends Lab Answers Free and Paid eBooks
  - Periodic Trends Lab Answers Public Domain eBooks
  - Periodic Trends Lab Answers eBook Subscription Services
  - Periodic Trends Lab Answers Budget-Friendly Options
- 6. Navigating Periodic Trends Lab Answers eBook Formats
  - o ePub, PDF, MOBI, and More
  - Periodic Trends Lab Answers Compatibility with Devices
  - Periodic Trends Lab Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Periodic Trends Lab Answers
  - Highlighting and Note-Taking Periodic Trends Lab Answers
  - Interactive Elements Periodic Trends Lab Answers
- 8. Staying Engaged with Periodic Trends Lab Answers
  - o Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Periodic Trends Lab Answers
- 9. Balancing eBooks and Physical Books Periodic Trends Lab Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Periodic Trends Lab Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Periodic Trends Lab Answers
  - Setting Reading Goals Periodic Trends Lab Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Periodic Trends Lab Answers
  - Fact-Checking eBook Content of Periodic Trends Lab Answers
  - Distinguishing Credible Sources

- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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