

PERIODIC TABLE (ADVANCED)

Molecular Structure of Hydrogen



Basic Information

Name: Hydrogen

Symbol: H

Atomic Number: 1

Atomic Mass: 1.00794 amu

Melting Point: -259.14°C ($-314.544495^{\circ}\text{K}$, $-434.41203^{\circ}\text{F}$)

Boiling Point: -252.87°C ($-335.385995^{\circ}\text{K}$, -423.166°F)

Number of Protons/Electrons: 1

Number of Neutrons: 0

Classification: Non-metal

Crystal Structure: Hexagonal

Density @ 293°K : 8.9898 g/cm^3

Color: colorless

Atomic Structure

Number of Energy Levels: 1

First Energy Level: 1

Isotopes

Isotope	Half Life
H-1	Stable
H-2	Stable
H-3	12.5 years

Facts

Name Origin: Greek words *hydro* (water) and *genes* (generated)

Uses: Balloons, metal refining

Obtained From: natural, all gas water

Molecular Structure of Carbon



Name: Carbon

Symbol: C

Atomic Number: 6

Atomic Mass: 12.0097 amu

Melting Point: 3500°C (3773.15°K , 6332.8°F)

Boiling Point: 4602.1°C (5100.15°K , 8320.6°F)

Number of Protons/Electrons: 6

Number of Neutrons: 6

Classification: Non-metal

Crystal Structure: Hexagonal

Density @ 293°K : 2.62 g/cm^3

Color: May be black

Atomic Structure

Number of Energy Levels: 2

First Energy Level: 2

Second Energy Level: 4

Isotopes

Isotope	Half Life
C-11	20.3 minutes
C-12	Stable
C-13	Stable
C-14	5730 years
C-15	2.3 seconds

Facts

Uses: steel, diamonds

Obtained From: burning with insufficient oxygen

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1 what is the difference between isotropic and anisotropic materials if a material exhibits same mechanical properties regardless of loading direction it is isotropic e g homogeneous cast materials materials lacking this property are anisotropic 2 what are orthotropic materials

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energy can only be transferred or changed from one form to another the second law of thermodynamics states that the total entropy of an isolated system can never reduce over time third law of thermodynamics states that if an object reaches the absolute zero temperature 0 k 273 15°c its atoms will stop moving

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