

GRAPHING PERIODIC TRENDS

Standard: Students know how to use the periodic table to identify trends in ionization energy, electronegativity, and the relative sizes of ions and atoms.

PRE-LAB DISCUSSION:

The Periodic Table is arranged according to the Periodic Law. The Periodic Law states that when elements are arranged in order of increasing atomic number, their physical and chemical properties show a periodic pattern. Students can discover these patterns by examining the changes in properties of elements on the Periodic Table. The properties that will be examined in this lesson are: atomic radius AND first ionization energy.

PURPOSE:

To understand periodic trends relating to atomic radius, ionization energy, and electronegativity.

PROCEDURE:

Use the information in these tables to complete the graph as described below.

RESULTS:

Symbol	Atomic Radius (Picometers)	First Ionization (Kilojoules/mole)	Electronegativity (4-point scale)
H	37	1312	2.1
Li	128	520	1.0
Na	190	496	0.9
K	203	419	0.8
Rb	220	403	0.8
Cs	244	376	0.7

Symbol	Atomic Radius (Picometers)	First Ionization (Kilojoules/mole)	Electronegativity (4-point scale)
He	31	498	0.0
Mg	141	738	1.2
Al	121	578	1.5
Sa	111	786	1.8
P	107	1012	2.1
S	106	1000	2.5
Cl	102	1251	3.0
Ar	100	1521	—

1. Define each of these terms: Atomic radius, First Ionization Energy, and Electronegativity.
2. Using colored-pencil or pen, list under the symbol (in this order) the Atomic Radius, First Ionization Energy, and Electronegativity. Use a different color for each property. Example: Write all of the Atomic radius values in red, all of the First Ionization energies in green, and all of the Electronegativities in blue. Of course, you can pick any colors that are available, as long as you are consistent.
3. Observe the trends in each property as you go down the Alkali metal group, and as you go across Period 3.
4. Write out each of the statements written on the opposite side, completing each statement with the observed trend (increase or decrease).
5. There is NO ABSTRACT to write for this activity :)

Periodic Properties Graphing Lab Answer Key

TD Snyder



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Table of Contents Periodic Properties Graphing Lab Answer Key

1. Understanding the eBook Periodic Properties Graphing Lab Answer Key
 - The Rise of Digital Reading Periodic Properties Graphing Lab Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Properties Graphing Lab Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Properties Graphing Lab Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Properties Graphing Lab Answer Key
 - Personalized Recommendations
 - Periodic Properties Graphing Lab Answer Key User Reviews and Ratings

- Periodic Properties Graphing Lab Answer Key and Bestseller Lists
- 5. Accessing Periodic Properties Graphing Lab Answer Key Free and Paid eBooks
 - Periodic Properties Graphing Lab Answer Key Public Domain eBooks
 - Periodic Properties Graphing Lab Answer Key eBook Subscription Services
 - Periodic Properties Graphing Lab Answer Key Budget-Friendly Options
- 6. Navigating Periodic Properties Graphing Lab Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Periodic Properties Graphing Lab Answer Key Compatibility with Devices
 - Periodic Properties Graphing Lab Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Properties Graphing Lab Answer Key
 - Highlighting and Note-Taking Periodic Properties Graphing Lab Answer Key
 - Interactive Elements Periodic Properties Graphing Lab Answer Key
- 8. Staying Engaged with Periodic Properties Graphing Lab Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Properties Graphing Lab Answer Key
- 9. Balancing eBooks and Physical Books Periodic Properties Graphing Lab Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Properties Graphing Lab Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Periodic Properties Graphing Lab Answer Key
 - Setting Reading Goals Periodic Properties Graphing Lab Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Periodic Properties Graphing Lab Answer Key
 - Fact-Checking eBook Content of Periodic Properties Graphing Lab Answer Key
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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