

POGIL: Periodic Table Trends

(Adapted from Rush Henrietta CSD)

Why:

The Periodic Table is one of the greatest inventions in the history of man. It allows scientists to predict physical and chemical properties of the elements. Dimitri Mendeleev (a Russian scientist) and Robert Mosley (a British chemist) put together this table in the late 1800's based on properties of the elements known at the time. The trends in these properties as you go across periods and down groups is the subject of this Chem POGIL.

Periodic Table of the Elements

Legend:

- Alkali metals
- Alkaline earth metals
- Transition metals
- Post-transition metals
- Nonmetals
- Metalloids
- Chalcogens
- Halogens
- Noble gases

Success Criteria:

- Understand the meaning of *atomic radius, reactivity, electronegativity, ionization energy*
- Recognize trends in atomic radius, reactivity, ionization energy and electronegativity as you go across periods and down groups.

Resources:

Reference Table S, Periodic Table, Unit 3 (Chapter 6) Notes, & Chapter 6 in textbook

Prerequisites:

Answer the following questions using the resources above.

- What happens to the number of valence electrons as you go down a group on the periodic table?
- What happens to the number of valence electrons as you go from left to right across a period?
- What happens to the number of energy levels as you go down a group on the periodic table?
- What happens to the number of energy levels as you go from left to right across a period?
- Draw a Lewis structure for a Cl atom.
- Draw a Bohr diagram for the Cl atom. (Draw above!)
- In this lab we will be talking about atomic radius.
 - Look up the atomic radius of Cl on Table S: _____
 - Why is it possible to show the atomic radius using a Bohr diagram but not by using a Lewis structure?
 - On your Bohr diagram draw an arrow to represent the radius of the Cl atom.
- Define *periodic*:
- Define *trend*:

Periodic Trends Answers

Lauren Gardner



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Table of Contents Periodic Trends Answers

1. Understanding the eBook Periodic Trends Answers
 - The Rise of Digital Reading Periodic Trends Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Trends Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Trends Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Trends Answers
 - Personalized Recommendations

- Periodic Trends Answers User Reviews and Ratings
- Periodic Trends Answers and Bestseller Lists
- 5. Accessing Periodic Trends Answers Free and Paid eBooks
 - Periodic Trends Answers Public Domain eBooks
 - Periodic Trends Answers eBook Subscription Services
 - Periodic Trends Answers Budget-Friendly Options
- 6. Navigating Periodic Trends Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Periodic Trends Answers Compatibility with Devices
 - Periodic Trends Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Trends Answers
 - Highlighting and Note-Taking Periodic Trends Answers
 - Interactive Elements Periodic Trends Answers
- 8. Staying Engaged with Periodic Trends Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Trends Answers
- 9. Balancing eBooks and Physical Books Periodic Trends Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Trends Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Periodic Trends Answers
 - Setting Reading Goals Periodic Trends Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Periodic Trends Answers
 - Fact-Checking eBook Content of Periodic Trends Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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