

## Trends on the Periodic Table

### Part 1: Identifying chemical properties/characteristics of the elements

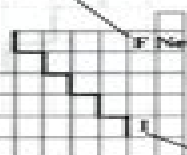
For each of the following elements, draw a caricature or cartoon above the description to show its "chemical personality" based on the description provided. (Note: See the Neon example for ideas.)

**Lithium**  
Being one of the smaller elements with only 3 electrons, lithium is happy to give one of its electrons away and never seems to ask for any in return.



**Francium**  
Francium is the largest of all of the elements and has so many electrons that it has trouble keeping track of them. Francium will eagerly give up electrons to other elements and never takes in extras. But watch out, it has an extremely reactive and explosive personality!

**Fluorine**  
One of the smallest elements, Fluorine is often described as snappy with a "magnetic" personality. As part of the halogen family (stoutest electron thieves), Fluorine is the best at convincing other elements to give up their electrons to him. However, he holds on tight to his own electrons and will not let them go without a fight.



I am so perfect... I have a full valence shell. I don't need those other elements... I hope they just leave me alone!



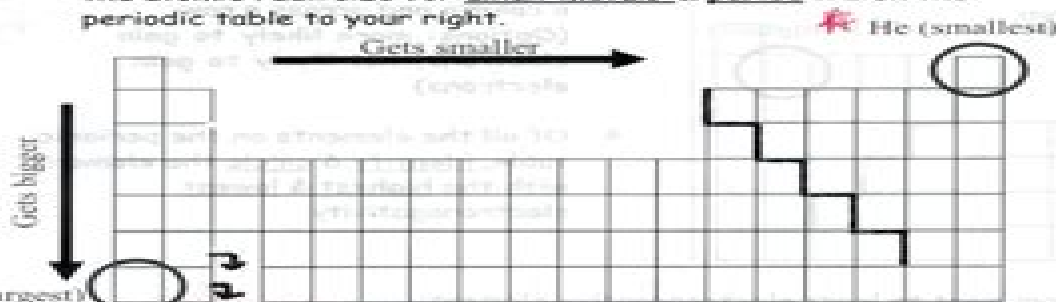
**Neon**  
Part of the Noble family, Neon does not readily interact with other elements. With a full set of valence electrons, Neon feels that he has everything he needs to make him happy without sharing with others.

**Iodine**  
As part of the halogen family, Iodine is also a skilled electron thief, however, not nearly as good as its smaller brothers and sisters. Iodine will also put up a pretty good fight to keep its own electrons.

### Part 2: Use your textbook to complete the following:

#### Atomic Radii Size:

1. Explain **atomic radius**: A measure of the size (volume) of an atom. Determined by measuring half of the distance between two identical atoms bonded together.
2. Using the information in figure 5-13 (pg 141) in your book, **write** the atomic radii size for **alkali metals** & **period #2** on the periodic table to your right.



3. What happens to atomic radii as you move down a column and across a row? On the arrows on the diagram, indicate the trend. (Options: gets bigger OR gets smaller)
4. Of all the elements on the periodic table, **identify** & **circle** the **largest** & **smallest** in size.

5. Rank the following elements from largest to smallest in atomic radii:

a) Be, Ba, Mg = **Ba, Mg, Be**

b) Li, F, B = **Li, B, F**

6. How can you explain the trend?

a. Why do they get smaller as you go across a period? Additional protons in the nucleus "pull" on electrons, bring the electron cloud closer.

*Effective nuclear charge,  $Z_{eff}$ , increases the attraction of the nucleus and therefore pulls the electron cloud closer to the nucleus resulting in smaller radii.*

*not a good enough answer*  
**4AP Answer**

# Periodic Table Answers Pogil

**CH Cherryholmes**



## **Periodic Table Answers Pogil:**

## **Periodic Table Answers Pogil** Book Review: Unveiling the Power of Words

In a global driven by information and connectivity, the energy of words has be more evident than ever. They have the capability to inspire, provoke, and ignite change. Such could be the essence of the book **Periodic Table Answers Pogil**, a literary masterpiece that delves deep to the significance of words and their impact on our lives. Compiled by a renowned author, this captivating work takes readers on a transformative journey, unraveling the secrets and potential behind every word. In this review, we will explore the book is key themes, examine its writing style, and analyze its overall impact on readers.

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