

## Calculus-Based Thermodynamics (Practice Problems)

1. What will be the freezing-point and boiling-point of an aqueous solution containing 100 g of glucose,  $C_6H_{12}O_6$ , and 100 g of water? Equations:  $K_f = 1.86^\circ C/m$ ,  $K_b = 0.512^\circ C/m$ .

$$\text{Molality} = \frac{\text{moles solute}}{\text{kg solvent}} = \frac{100 \text{ g} \times \frac{1 \text{ mol}}{180 \text{ g}}}{0.100 \text{ kg}} = 5.56 \text{ mol/kg}$$

$$\Delta T_f = K_f \cdot m = 1.86^\circ C/m \cdot 5.56 \text{ mol/kg} = 10.3^\circ C \quad \Delta T_b = K_b \cdot m = 0.512^\circ C/m \cdot 5.56 \text{ mol/kg}$$

2. What will be the freezing-point and boiling-point of an aqueous solution containing 100 g of NaCl and 100 g of water? Equations:  $K_f = 1.86^\circ C/m$ ,  $K_b = 0.512^\circ C/m$ ,  $K_f(\text{NaCl}) = 3.72^\circ C/m$ ,  $K_b(\text{NaCl}) = 1.02^\circ C/m$ .

$$\text{Molality} = \frac{\text{moles solute}}{\text{kg solvent}} = \frac{100 \text{ g} \times \frac{1 \text{ mol}}{58.4 \text{ g}}}{0.100 \text{ kg}} = 1.71 \text{ mol/kg}$$

$$\Delta T_f = K_f(\text{NaCl}) \cdot m = 3.72^\circ C/m \cdot 1.71 \text{ mol/kg} = 6.36^\circ C \quad \Delta T_b = K_b(\text{NaCl}) \cdot m = 1.02^\circ C/m \cdot 1.71 \text{ mol/kg}$$

3. A 100 mL sample of 0.10 M  $NaCl$  is completely mixed with 100 g of water. Calculate the freezing-point depression.  $K_f = 1.86^\circ C/m$ . Assume that the solute does not undergo any further decomposition.

$$\text{Molality} = \frac{\text{moles solute}}{\text{kg solvent}} = \frac{0.10 \text{ mol}}{0.100 \text{ kg}} = 1.00 \text{ mol/kg}$$

$$\Delta T_f = K_f \cdot m = 1.86^\circ C/m \cdot 1.00 \text{ mol/kg}$$

$$\Delta T_f = 1.86^\circ C$$

$$\Delta T_f = 1.86^\circ C$$

$$\Delta T_f = K_f(\text{NaCl}) \cdot m = 3.72^\circ C/m \cdot 1.00 \text{ mol/kg}$$

$$\Delta T_f = 3.72^\circ C$$

$$\Delta T_f = 3.72^\circ C$$

$$\Delta T_f = K_f(\text{NaCl}) \cdot m = 3.72^\circ C/m \cdot 1.00 \text{ mol/kg} = 3.72^\circ C$$

4. A solution composed from 10 g of glucose,  $C_6H_{12}O_6$ , and 100 g of water is prepared. The freezing-point of this solution is  $-1.7^\circ C$ . The normal freezing point of pure water is  $0^\circ C$  and the freezing-point depression constant for water is  $1.86^\circ C/m$ . Determine the molar mass of the compound. Assume that the solute does not undergo any further decomposition.

$$\Delta T_f = K_f \cdot m \quad 1.7^\circ C = 1.86^\circ C/m \cdot m \quad m = 0.914 \text{ mol/kg}$$

$$\text{Molality} = \frac{\text{moles solute}}{\text{kg solvent}}$$

$$0.914 \text{ mol/kg} = \frac{10 \text{ g}}{100 \text{ g}} \cdot \frac{1 \text{ mol}}{M}$$

$$M = 10.9 \text{ g/mol}$$

$$\Delta T_f = K_f(\text{NaCl}) \cdot m = 3.72^\circ C/m \cdot 0.914 \text{ mol/kg}$$

$$\Delta T_f = 3.4^\circ C$$

$$\Delta T_f = 3.4^\circ C$$

$$\Delta T_f = K_f \cdot m \quad 1.7^\circ C = 1.86^\circ C/m \cdot m \quad m = 0.914 \text{ mol/kg}$$

$$\Delta T_f = 3.4^\circ C$$

# Rosengarten Colligative Properties Homework Answers

**Peipei Pang**



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