

## **PRE-AP CHEMISTRY: Semester 2 Final Review**

**Final: June 8 (period 7); June 9 (period 6)**

Review days after school: Monday June 1, Wednesday June 3, Thursday June 4

\*You have come a long way this year! You should celebrate your successes and continue to work hard until the end. Let me know how I can help you!

\*\*You can also use a periodic table and your reference sheets from the semester.

\*\*\*Your binder contains your work from the whole semester (maybe even the whole year!) You can study from your semester's old worksheets, practice problems within the chapters, labs, quizzes, or tests, or look in my folders for extra worksheets/practice.

### **I. SEMESTER 1:**

- a. Significant figures: rules for measuring, rounding, etc.
  - i. Why do we need them?
  - ii. How can you tell how many sig figs are in a measurement?
- b. Scientific notation, converting metric units
  - i.  $5.64 \times 10^3 =$  \_\_\_\_\_
  - ii.  $0.000374 =$  \_\_\_\_\_
  - iii.  $5000 \text{ grams} =$  \_\_\_\_\_ kg
  - iv.  $0.345 \text{ L} =$  \_\_\_\_\_ mL
- c. Physical/chemical changes (in Ch. 9 again)
- d. Structure of an atom: numbers of protons, neutrons, electrons
  - i. What are isotopes?
  - ii. How do you identify the number of protons, neutrons, and electrons in an atom?
- e. Electron configurations
  - i. Use the Aufbau principle to write  $e^-$  configs for any element on P-table
  - ii. Know electron configurations for **ions**; what has changed? Why do atoms become ions?
- f. Arrangement of periodic table
  - i. Names of specific groups on table: halogens, alkali metals, noble gases, metals, nonmetals

### **II. UNIT 5 IONIC BONDING (Chapter 4/12)**

- i. How do positive ions form? Negative ions?
- ii. What is the octet rule? How does ionic bonding happen?
- iii. Naming and writing formulas, including those with transition metals
  1.  $\text{MgCl}_2 =$  \_\_\_\_\_
  2.  $\text{KF} =$  \_\_\_\_\_
  3. Iron (III) oxide = \_\_\_\_\_
- iv. Common polyatomic ions
  1.  $\text{Na}_2\text{CO}_3 =$  \_\_\_\_\_
  2. Aluminum hydroxide = \_\_\_\_\_

### **III. UNIT 5: COVALENT BONDING**

- a. What is the octet rule? How does covalent bonding happen?
- b. Single, double, triple bonds + valence electrons
- c. Naming and writing formulas (system of prefixes)
- d. VSEPR: predicting shapes of molecules
- e. Polar/nonpolar: what is it, how do you know?
  - i. Draw Lewis structures for the following molecules and identify the shape and polarity:
    1.  $\text{H}_2\text{O}$
    2.  $\text{NH}_3$
    3.  $\text{CH}_4$
    4.  $\text{SF}_6$
    5.  $\text{SO}_4^{2-}$

### **IV. UNIT 6/7: CHEMICAL REACTIONS**

- a. What is a chemical reaction?

# Pre Ap English 8 Semester 1 Exam Review

**Richard Bailey**



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## **Table of Contents Pre Ap English 8 Semester 1 Exam Review**

1. Understanding the eBook Pre Ap English 8 Semester 1 Exam Review
  - The Rise of Digital Reading Pre Ap English 8 Semester 1 Exam Review
  - Advantages of eBooks Over Traditional Books
2. Identifying Pre Ap English 8 Semester 1 Exam Review
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Pre Ap English 8 Semester 1 Exam Review
  - User-Friendly Interface
4. Exploring eBook Recommendations from Pre Ap English 8 Semester 1 Exam Review
  - Personalized Recommendations
  - Pre Ap English 8 Semester 1 Exam Review User Reviews and Ratings
  - Pre Ap English 8 Semester 1 Exam Review and Bestseller Lists
5. Accessing Pre Ap English 8 Semester 1 Exam Review Free and Paid eBooks

- Pre Ap English 8 Semester 1 Exam Review Public Domain eBooks
- Pre Ap English 8 Semester 1 Exam Review eBook Subscription Services
- Pre Ap English 8 Semester 1 Exam Review Budget-Friendly Options
- 6. Navigating Pre Ap English 8 Semester 1 Exam Review eBook Formats
  - ePub, PDF, MOBI, and More
  - Pre Ap English 8 Semester 1 Exam Review Compatibility with Devices
  - Pre Ap English 8 Semester 1 Exam Review Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Pre Ap English 8 Semester 1 Exam Review
  - Highlighting and Note-Taking Pre Ap English 8 Semester 1 Exam Review
  - Interactive Elements Pre Ap English 8 Semester 1 Exam Review
- 8. Staying Engaged with Pre Ap English 8 Semester 1 Exam Review
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Pre Ap English 8 Semester 1 Exam Review
- 9. Balancing eBooks and Physical Books Pre Ap English 8 Semester 1 Exam Review
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Pre Ap English 8 Semester 1 Exam Review
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Pre Ap English 8 Semester 1 Exam Review
  - Setting Reading Goals Pre Ap English 8 Semester 1 Exam Review
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Pre Ap English 8 Semester 1 Exam Review
  - Fact-Checking eBook Content of Pre Ap English 8 Semester 1 Exam Review
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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web reading the middle predicting the beginning and the end while reading tasks although reading is often a solitary activity and the idea of reading in pairs seems odd reading can be collaborative approaches i use include

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web during lessons students must be involved in activities which enable them to respond cognitively emotionally and imaginatively to imaginative writing several of the activities found in this section can be adapted for use in the section ideas for after the reading and vice versa activity 1

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