



COVALENT BONDING

Practice Problems

Do your answers to these very practice problems.

SECTION 8.1 MOLECULAR COMPOUNDS

1. Classify each of the following compounds as ionic or covalent.

- a. Na_2O (ionic) b. CH_4 (covalent) c. HCl (covalent)
 d. KNO_3 (ionic) e. H_2O (covalent)

2. Which of the following are molecular compounds?

- a. H_2O b. NaCl c. CO_2
 d. KCl e. H_2SO_4

3. Which of the following are molecular compounds? Write the formula for each.

a. Carbon dioxide is a CO_2 molecular compound. The carbon atom is bonded to two oxygen atoms, and each oxygen atom is bonded to one carbon atom.

b. Sulfur dioxide is a SO_2 molecular compound. The sulfur atom is bonded to two oxygen atoms, and each oxygen atom is bonded to one sulfur atom.

SECTION 8.2 THE NATURE OF COVALENT BONDING

1. Draw the Lewis structure for hydrogen fluoride, HF .



2. Draw the Lewis structure for methane, CH_4 .



3. Draw the Lewis structure for carbon dioxide, CO_2 .



4. Draw the Lewis structure for acetylene, C_2H_2 .

5. Draw the Lewis structure for hydrogen cyanide, HCN . Do this by drawing the structure for HNC .



Answer key:

SECTION 8.3 BONDING: THE ORBITAL

1. Which orbital has two lobes, one on each side of the nucleus? p orbital. The lobes contain the two electrons.

2. Which orbital has two lobes, one on each side of the nucleus? p orbital. The lobes contain the two electrons.

Practice Problems Chapter 8 Covalent Bonding Worksheet

A Loxley



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