

# Periodic Trends Practice



## Atomic Radius

The distance from the nucleus of one bonded atom.



## Electronegativity

How much an atom "wants" electrons or the atom's tendency to attract electrons.



## Ionization Energy

The amount of energy required to remove electrons from an atom.



## Electron Affinity

The amount of energy released when a neutral gaseous atom gains an electron.



## Driving Question

How can the periodic table be used to predict atomic characteristics such as atomic radius, electronegativity, ionization energy, and electron affinity?

Describe what causes this trend...

As you move from left to right across a period, the number of protons increases, which increases the pull of the nucleus on the electrons, causing the atomic radius to decrease.

As you move from top to bottom down a group, the number of protons increases, but the number of electron shells also increases, which increases the distance between the nucleus and the electrons, causing the atomic radius to increase.

Describe what causes this trend...

As you move from left to right across a period, the number of protons increases, which increases the pull of the nucleus on the electrons, causing the electronegativity to increase.

As you move from top to bottom down a group, the number of protons increases, but the number of electron shells also increases, which increases the distance between the nucleus and the electrons, causing the electronegativity to decrease.

Describe what causes this trend...

As you move from left to right across a period, the number of protons increases, which increases the pull of the nucleus on the electrons, causing the ionization energy to increase.

As you move from top to bottom down a group, the number of protons increases, but the number of electron shells also increases, which increases the distance between the nucleus and the electrons, causing the ionization energy to decrease.

Describe what causes this trend...

As you move from left to right across a period, the number of protons increases, which increases the pull of the nucleus on the electrons, causing the electron affinity to increase.

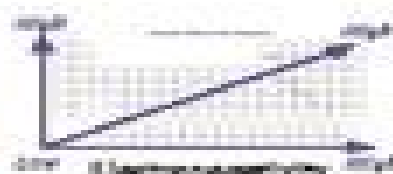
As you move from top to bottom down a group, the number of protons increases, but the number of electron shells also increases, which increases the distance between the nucleus and the electrons, causing the electron affinity to decrease.

## Shielding?

What is shielding? What effect does it have on the periodic trends defined above?

Shielding occurs when inner shell electrons block out the positive pull of the nucleus on the outer shell electrons. The more shielding there is, the more the trends of atomic radius, electronegativity, ionization energy, and electron affinity are affected.

Add arrows to each of the following periodic tables to indicate the direction of the trend as it increases or decreases from left to right and top to bottom. (Tip: Use a different color for each of the periodic tables so you can easily remember them.)



# Practice Identifying Periodic Trends

**Oswaal Editorial Board**



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## **Table of Contents Practice Identifying Periodic Trends**

1. Understanding the eBook Practice Identifying Periodic Trends
  - The Rise of Digital Reading Practice Identifying Periodic Trends
  - Advantages of eBooks Over Traditional Books
2. Identifying Practice Identifying Periodic Trends
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Practice Identifying Periodic Trends
  - User-Friendly Interface
4. Exploring eBook Recommendations from Practice Identifying Periodic Trends
  - Personalized Recommendations
  - Practice Identifying Periodic Trends User Reviews and Ratings
  - Practice Identifying Periodic Trends and Bestseller Lists

5. Accessing Practice Identifying Periodic Trends Free and Paid eBooks
  - Practice Identifying Periodic Trends Public Domain eBooks
  - Practice Identifying Periodic Trends eBook Subscription Services
  - Practice Identifying Periodic Trends Budget-Friendly Options
6. Navigating Practice Identifying Periodic Trends eBook Formats
  - ePub, PDF, MOBI, and More
  - Practice Identifying Periodic Trends Compatibility with Devices
  - Practice Identifying Periodic Trends Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Practice Identifying Periodic Trends
  - Highlighting and Note-Taking Practice Identifying Periodic Trends
  - Interactive Elements Practice Identifying Periodic Trends
8. Staying Engaged with Practice Identifying Periodic Trends
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Practice Identifying Periodic Trends
9. Balancing eBooks and Physical Books Practice Identifying Periodic Trends
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Practice Identifying Periodic Trends
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Practice Identifying Periodic Trends
  - Setting Reading Goals Practice Identifying Periodic Trends
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Practice Identifying Periodic Trends
  - Fact-Checking eBook Content of Practice Identifying Periodic Trends
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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