

## Practice Problems: Quantum Theory

CHEM 1A

1. Hydrogen has a red emission line at 656.3 nm, what is the energy and frequency of a photon of this light? Note: Planck's constant =  $6.626 \times 10^{-34}$  J·s, speed of light =  $2.998 \times 10^8$  m/s

$$E_{\text{photon}} = h\nu = \frac{hc}{\lambda}$$

$$E_{\text{photon}} = \frac{(6.626 \times 10^{-34} \text{ J}\cdot\text{s}) (2.998 \times 10^8 \text{ m/s})}{656.3 \text{ nm} \left( \frac{10^{-9} \text{ m}}{1 \text{ nm}} \right)}$$

$$E_{\text{photon}} = \underline{3.02677861 \times 10^{-19} \text{ J}}$$

$$\nu = \frac{c}{\lambda}$$

$$\nu = \frac{(2.998 \times 10^8 \text{ m/s})}{656.3 \text{ nm} \left( \frac{10^{-9} \text{ m}}{1 \text{ nm}} \right)}$$

$$\nu = \underline{4.568032912 \times 10^{14} \text{ s}^{-1}}$$

Answer (E):  $3.027 \times 10^{-19} \text{ J}$ Answer ( $\nu$ ):  $4.568 \times 10^{14} \text{ s}^{-1}$  or Hz

2. What is the de Broglie wavelength for an electron moving at  $2.42 \times 10^6$  m/s?  
Note:  $m_e = 9.109 \times 10^{-31}$  kg

$$\lambda = \left( \frac{h}{mv} \right)$$

$$\lambda = \left( \frac{6.626 \times 10^{-34} \text{ J}\cdot\text{s}}{(9.109 \times 10^{-31} \text{ kg}) (2.42 \times 10^6 \text{ m/s})} \right) = \underline{3.00583657 \times 10^{-10} \text{ m}}$$

Answer:  $3.01 \times 10^{-10} \text{ m}$   $\longleftarrow$  many orders of magnitude larger than the diameter of an electron

3. Using the Bohr equation, calculate is the change in energy when an electron transitions from  $n = 3$  to  $n = 2$  in a hydrogen atom.

$$\Delta E = -2.18 \times 10^{-18} \text{ J} \left( \frac{Z^2}{n_{\text{final}}^2} - \frac{Z^2}{n_{\text{initial}}^2} \right)$$

$$\Delta E = -2.18 \times 10^{-18} \text{ J} \left( \frac{1^2}{2^2} - \frac{1^2}{3^2} \right)$$

$$\Delta E = \underline{-3.02777778 \times 10^{-19} \text{ J}}$$

Answer:  $-3.03 \times 10^{-19} \text{ J}$

# Practice Problems Chapter 7 Quantum Theory

**Jörg C. Lemm**



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## **Table of Contents Practice Problems Chapter 7 Quantum Theory**

1. Understanding the eBook Practice Problems Chapter 7 Quantum Theory
  - The Rise of Digital Reading Practice Problems Chapter 7 Quantum Theory
  - Advantages of eBooks Over Traditional Books
2. Identifying Practice Problems Chapter 7 Quantum Theory
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Practice Problems Chapter 7 Quantum Theory
  - User-Friendly Interface
4. Exploring eBook Recommendations from Practice Problems Chapter 7 Quantum Theory
  - Personalized Recommendations
  - Practice Problems Chapter 7 Quantum Theory User Reviews and Ratings
  - Practice Problems Chapter 7 Quantum Theory and Bestseller Lists
5. Accessing Practice Problems Chapter 7 Quantum Theory Free and Paid eBooks
  - Practice Problems Chapter 7 Quantum Theory Public Domain eBooks
  - Practice Problems Chapter 7 Quantum Theory eBook Subscription Services
  - Practice Problems Chapter 7 Quantum Theory Budget-Friendly Options
6. Navigating Practice Problems Chapter 7 Quantum Theory eBook Formats

- ePub, PDF, MOBI, and More
  - Practice Problems Chapter 7 Quantum Theory Compatibility with Devices
  - Practice Problems Chapter 7 Quantum Theory Enhanced eBook Features
7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Practice Problems Chapter 7 Quantum Theory
    - Highlighting and Note-Taking Practice Problems Chapter 7 Quantum Theory
    - Interactive Elements Practice Problems Chapter 7 Quantum Theory
  8. Staying Engaged with Practice Problems Chapter 7 Quantum Theory
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Practice Problems Chapter 7 Quantum Theory
  9. Balancing eBooks and Physical Books Practice Problems Chapter 7 Quantum Theory
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Practice Problems Chapter 7 Quantum Theory
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Practice Problems Chapter 7 Quantum Theory
    - Setting Reading Goals Practice Problems Chapter 7 Quantum Theory
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Practice Problems Chapter 7 Quantum Theory
    - Fact-Checking eBook Content of Practice Problems Chapter 7 Quantum Theory
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development
    - Exploring Educational eBooks
  14. Embracing eBook Trends
    - Integration of Multimedia Elements
    - Interactive and Gamified eBooks

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